

14

THE BEHAVIOR OF GASES

SECTION 14.1 PROPERTIES OF GASES (pages 413–417)

This section uses kinetic theory to explain the properties of gases. This section also explains how gas pressure is affected by the amount of gas, its volume, and its temperature.

► Compressibility (pages 413–414)

1. Look at Figure 14.1 on page 413. How does an automobile air bag protect the crash dummy from being broken as a result of impact?

The gases used to inflate the airbag are able to absorb a considerable amount of energy when they are compressed.

2. What theory explains the behavior of gases? kinetic theory

3. Circle the letter next to each sentence that is true concerning the compressibility of gases.

- ☒ a. The large relative distances between particles in a gas means that there is considerable empty space between the particles.
- ☒ b. The assumption that particles in a gas are relatively far apart explains gas compressibility.
- ☒ c. Compressibility is a measure of how much the volume of matter decreases under pressure.
- d. Energy is released by a gas when it is compressed.

► Factors Affecting Gas Pressure (pages 414–417)

4. List the name, the symbol, and a common unit for the four variables that are generally used to describe the characteristics of a gas.

- a. pressure, P , kilopascals
- b. volume, V , liters
- c. temperature, T , kelvins
- d. amount of gas, n , moles

5. What keeps the raft in Figure 14.3 inflated?

The air pressure exerted by the enclosed gas keeps the raft inflated.

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