

# Percent Composition of Hydrates Investigation

[Name, Period, Rooming, Initial (C, L, R)]



## Background:

In this lab, we are looking at hydrated salt in which water molecules (spherical molecules that form a crystal lattice) are trapped within the lattice created by the  $\text{Cu}^{2+}$  cations and the  $\text{SO}_4^{2-}$  anions. This lattice is composed of both positive and negative charges. The positive  $\text{Cu}^{2+}$  ions are attracted to  $\text{OH}_2$  molecules. Water can be held this way between the layers of the crystal lattice structure. This is called hydrated copper sulfate (blue) or anhydrous.

Hydrated salts, like hydrated copper sulfate, can be written as  $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ . The "x" represents the number of water molecules associated with each unit of salt, as shown in anhydrous.

The objective is to determine the number water molecules associated with each unit of  $\text{CuSO}_4$ . Find "x" in  $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ .

Upon heating the hydrate (anhydrous) until all water has been removed,  $\text{CuSO}_4$  is produced. This balanced equation illustrates:



However, this is done with heat, plus the word "heat" is a delta symbol  $\Delta$  as in the photo above.

**Objective:** The objective is to determine the number water molecules associated with each unit of  $\text{CuSO}_4$ . Find "x" in  $\text{CuSO}_4 \cdot x\text{H}_2\text{O}$ .

**Claim (Hypothesis):** The number water molecules associated with each unit of  $\text{CuSO}_4$  is

**Materials:** Paper, glass, crucible, hydrated salt, ring stand, Bunsen burner, crucible tongs, balance, burner, and flask.

## Procedure:

1. Label each crucible in CL, R, A, M and a completely empty.
2. Weigh the crucible separately and then hydrated salt.
3. Place the crucible on the paper above the Bunsen burner and heat the crucible until the paper is white. Record the weight.
4. Weigh the hydrated salt in the crucible.
5. Heat until the  $\text{CuSO}_4$  and  $\text{H}_2\text{O}$  left by the crucible.
6. Repeat the weight of the crucible after the paper is white.

# Composition Of Hydrates Lab Report

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