

# REPORT SHEET

## Compounds and Their Bonds

LAB

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### A. Ions: Transfer of Electrons

Element	Atomic #	1. Electron Configuration of Atom	2. Electron Dot Symbol	3. Loss or Gain	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	Na <sup>•</sup>	lose 1 e <sup>-</sup>	$1s^2 2s^2 2p^6$	1 <sup>+</sup>	Na <sup>+</sup>	Sodium ion
Nitrogen	7	$1s^2 2s^2 2p^3$	• N •	Gain 3 e <sup>-</sup>	$1s^2 2s^2 2p^6$	3 <sup>-</sup>	N <sup>3-</sup>	Nitride ion
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	• Al •	Lose 3 e <sup>-</sup>	$1s^2 2s^2 2p^6$	3 <sup>+</sup>	Al <sup>3+</sup>	Aluminum ion
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$	: Cl •	Gain 1 e <sup>-</sup>	$1s^2 2s^2 2p^6 3s^2 3p^6$	1 <sup>-</sup>	Cl <sup>-</sup>	Chloride ion
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$	Ca <sup>••</sup>	Lose 2 e <sup>-</sup>	$1s^2 2s^2 2p^6 3s^2 3p^4$	2 <sup>+</sup>	Ca <sup>2+</sup>	Calcium ion
Oxygen	8	$1s^2 2s^2 2p^4$	• O •	Gain 2 e <sup>-</sup>	$1s^2 2s^2 2p^6$	2 <sup>-</sup>	O <sup>2-</sup>	Oxide ion

### B. Ionic Compounds and Formulas

#### 1. Physical properties

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	White, solid, crystalline	2.16 g/cm <sup>3</sup>	801°C

#### 2. Formulas of ionic compounds

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li <sup>1+</sup>	I <sup>1-</sup>	LiI
Aluminum oxide	Al <sup>3+</sup>	O <sup>2-</sup>	Al <sub>2</sub> O <sub>3</sub>
Calcium sulfide	Ca <sup>2+</sup>	S <sup>2-</sup>	CaS
Magnesium bromide	Mg <sup>2+</sup>	Br <sup>1-</sup>	MgBr <sub>2</sub>
Potassium nitride	K <sup>1+</sup>	N <sup>3-</sup>	K <sub>3</sub> N
Sodium fluoride	Na <sup>1+</sup>	F <sup>1-</sup>	NaF

#### 3. Names of ionic compounds

Formula	Positive Ion	Negative Ion	Name
K <sub>2</sub> S	K <sup>1+</sup>	S <sup>2-</sup>	Potassium sulfide
BaF <sub>2</sub>	Ba <sup>2+</sup>	F <sup>1-</sup>	Barium fluoride
MgO	Mg <sup>2+</sup>	O <sup>2-</sup>	Magnesium oxide
Na <sub>3</sub> N	Na <sup>1+</sup>	N <sup>3-</sup>	Sodium nitride
AlCl <sub>3</sub>	Al <sup>3+</sup>	Cl <sup>1-</sup>	Aluminum chloride
Mg <sub>3</sub> P <sub>2</sub>	Mg <sup>2+</sup>	P <sup>3-</sup>	Magnesium phosphide

# Compounds Their Formulas Lab Answers

**James T. Shipman, Jerry D.  
Wilson, Aaron W. Todd**



## Compounds Their Formulas Lab Answers:

*Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **Lab Manual for General,**

**Organic, and Biochemistry** Denise Guinn, Rebecca Brewer, 2009-08-21 Teaching all of the necessary concepts within the constraints of a one term chemistry course can be challenging Authors Denise Guinn and Rebecca Brewer have drawn on their 14 years of experience with the one term course to write a textbook that incorporates biochemistry and organic chemistry throughout each chapter emphasizes cases related to allied health and provides students with the practical quantitative skills they will need in their professional lives *Essentials of General Organic and Biochemistry* captures student interest from day one with a focus on attention getting applications relevant to health care professionals and as much pertinent chemistry as is reasonably possible in a one term course Students value their experience with chemistry getting a true sense of just how relevant it is to their chosen profession To browse a sample chapter view sample ChemCasts and more visit [www.whfreeman.com](http://www.whfreeman.com) **Lab Experiments in Introductory Chemistry** Phil Reedy, Donald J. Wink, Sharon

Fetzer-Gislason, 2003-03-21 The manual contains laboratory experiments written specifically for the prep chem lab as well as for the general chemistry course Available as a complete manual or custom published at [http://custompub.whfreeman.com](http://http://custompub.whfreeman.com)

**Spectroscopic Methods Lab** Mr. Rohit Manglik, 2024-05-18 EduGorilla Publication is a trusted name in the education sector committed to empowering learners with high quality study materials and resources Specializing in competitive exams and academic support EduGorilla provides comprehensive and well structured content tailored to meet the needs of students across various streams and levels **Geology From Experience** E. Kirsten Peters, Larry E. Davis, 2000-11-05 Moving away

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Timberlake, 2002-06-24 Drawing from the successful main Laboratory Manual the *Essential Laboratory Manual* includes twenty one experiments which have been revised and updated Suitable for a one or two term lab course **U.S.**

**Government Research & Development Reports** ,1970 *Chemical Abstracts* ,1909 **Chambers's Encyclopædia:**

**Goo.-Lab** ,1880 *Dept. Bulletins [on Forestry]* United States. Department of Agriculture, 1920 *Bulletin of the U.S.*

*Department of Agriculture* United States. Department of Agriculture, 1923 *Pamphlets on Forest Protection* ,1916

**Bibliography of Scientific and Industrial Reports** ,1971      **Technical Information Pilot** ,1949      *Journal of the American Pharmaceutical Association* American Pharmaceutical Association,1929      *American Chemical Journal* Ira Remsen,Charles August Rouiller,1900      *The National Druggist* ,1927      **The Coal Trade Bulletin** ,1912      Physiological Abstracts ,1922      **An Introduction to Physical Science** James T. Shipman,Jerry D. Wilson,Aaron W. Todd,2002-07-08

This text presents a survey of the physical sciences physics chemistry astronomy meteorology and geology for non science majors Topics are treated both descriptively and quantitatively providing flexibility for instructors who wish to emphasize a highly descriptive approach a highly quantitative approach or any spectrum in between

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