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#### CBSE Boards 2023

Answer & Solutions

#### Chemistry

#### Section-A

- Which one of the following has lowest pK, value?
  - A.  $CH_3 COOH$
  - B.  $O_2N CN_2 COOH$
  - C.  $C\hat{I} CH_2 COOH$
  - D. MCOOK

#### Answer (B)

#### Scal.

As,  $pK_n = -\log K_n$ 

So, as pK<sub>a</sub> decreases, acidity will increase, i.e. pK<sub>a</sub> a acidic strength a -1 character

Among the given options,  $-NO_2$  has highest -I character i.e.,  $O_2N - CH_2 - COOH$  has highest acidic character i.e. it will have lowest  $pK_a$  value.

Hence, option (b) is correct answer.

- BRUUS Which one of the following cell was used in Apollo space program?
  - A. Mercury cell
  - B. MRII
  - C. H<sub>2</sub> = O<sub>2</sub> Fuel cell
  - D. Dry cell

#### Answer (C)

#### Sol.

One of the most successful fuel cell uses the reaction of hydrogen with oxygen to form water. This fuel cell was used for providing electric power in Apollo space program.

Consider the following standard electrode potential values:

$$Fe_{aq}^{3+} + e^- \rightarrow Fe_{aq}^{2+}$$
;  $E^- = +0.77 V$ 

 $MnO_{1 \text{ os}}^{-} + 8H^{+} + 5e^{-} \rightarrow Mn_{ag}^{2+} + 4H_{2}O_{1}$ ;  $E^{a} = +1.51 \text{ V}$ 

What is the cell potential for the redox reaction?

- A. -2.28 V
- B. -0.74 V
- C. +0.74 F
- D. +2.28 V

#### Answer (C)

#### South

$$Fe_{\alpha q}^{3+} + e^- \rightarrow Fe_{\alpha q}^{2+}$$
;  $E^- = +0.77 V$   
 $MnO_{4-\alpha q}^- + 8H^+ + 5e^- \rightarrow Mn^{2+} + 4H_2O$ ;  $E^- = +1.51 V$ 

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