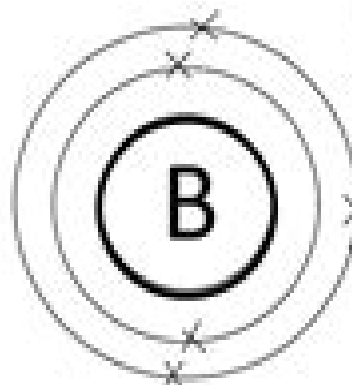


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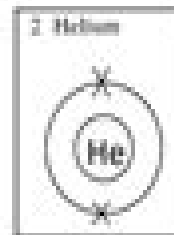
1 Hydrogen



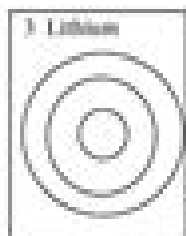
The X is an electron.

The circle O is a shell that electrons sit on.

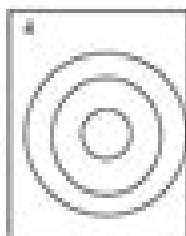
2 Helium



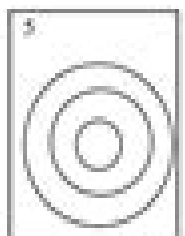
3 Lithium



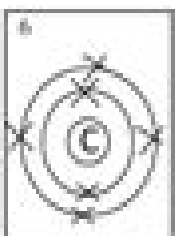
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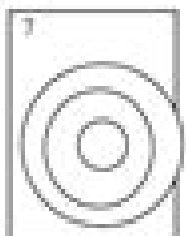
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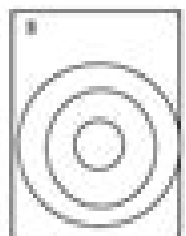
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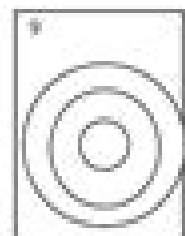
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9



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11 Sodium



12



13



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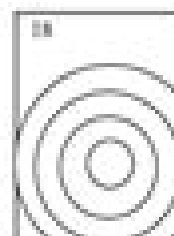
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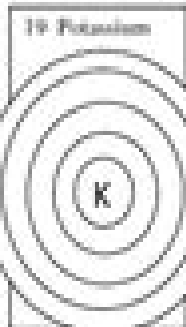
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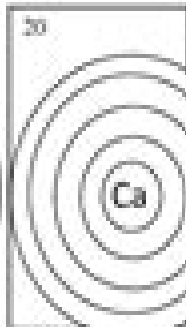
18



19 Potassium



20



Fill in each box with its electron shell pattern.

Remember:

- 1) The 1st shell can have up to **2** electrons
- 2) The 2nd shell can have up to **8** electrons
- 3) The 3rd shell can have up to **8** electrons

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Mark Steyvers

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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spontaneously it is in essence an attribute of individual atomic nuclei radioactive decay is a property of several naturally occurring elements as well as of artificially produced isotopes of the elements

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radioactivity some nuclides have an atomic nucleus that decays spontaneously without any external influence known as radionuclides they transform themselves into other nuclei ionising radiation is emitted during this process this is known as radioactivity

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it some future development of atomic theory perhaps a modification of the thomson model would most likely yield a causal explanation of radioactivity in terms of the internal structure of the atom

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