

Bond Enthalpy (Bond Energy)

Bond enthalpy, also known as **bond energy**, is defined as the amount of energy required to break a bond. These values are positive, indicating that bond breaking requires energy and is endothermic. Bond enthalpies are reported in kilojoules per mol (kJ/mol). The average energy for breaking a hydrogen-hydrogen bond is 436 kJ/mol.

In a chemical reaction several bonds are broken and formed. For example, in the reaction below a hydrogen-hydrogen bond is broken and a fluorine-fluorine bond is broken. Two hydrogen-fluorine bonds are formed. The overall enthalpy change for this reaction is calculated below:



Bond Energy: 436 kJ/mol



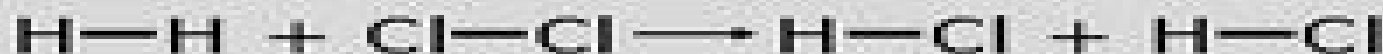
Bond Energy: 155 kJ/mol



Bond Energy: 567 kJ/mol

$$\begin{aligned} \Delta H_{\text{rxn}} &= [\text{energy used for breaking bonds}] - [\text{energy formed in making bonds}] \\ &= [436 \text{ kJ/mol} + 155 \text{ kJ/mol}] - [2(567 \text{ kJ/mol})] \\ &= -543 \text{ kJ/mol} \end{aligned}$$

Estimate the enthalpy of reaction (ΔH_{rxn}) using bond energy values.



Work area



Work area

Bond Energy Key

Grégory Guisbiers, Dibyendu Ganguli

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the authors' names.

Bond Energy Key:

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