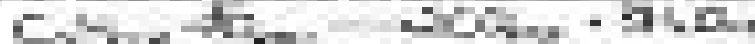


B = Bonding & Molecular Structure

Table B.3 • Some Average Single and Multiple Bond Energies (kJ/mol)

	H	Li	Na	K	Rb	Cs	F	Cl	Br	I	At
H-H	436										
Li-H		192									
Na-H			187								
K-H				164							
Rb-H					147						
Cs-H						140					
F-F							158				
Cl-Cl								242			
Br-Br									193		
I-I										151	
At-At											100
C-C							347				
C-H							413				
C-Cl								339			
C-Br									276		
C-I										238	
C-At											222
C≡C							837				
C≡N							891				
C≡O							1072				
C≡N ₂								812			
C≡O ₂									602		
C≡O ₃										527	
C≡O ₄											470
C≡O ₅											423
C≡O ₆											386
C≡O ₇											349
C≡O ₈											312
C≡O ₉											275
C≡O ₁₀											238
C≡O ₁₁											201
C≡O ₁₂											164
C≡O ₁₃											127
C≡O ₁₄											90
C≡O ₁₅											53
C≡O ₁₆											16
C≡O ₁₇											-21
C≡O ₁₈											-58
C≡O ₁₉											-95
C≡O ₂₀											-132
C≡O ₂₁											-169
C≡O ₂₂											-206
C≡O ₂₃											-243
C≡O ₂₄											-280
C≡O ₂₅											-317
C≡O ₂₆											-354
C≡O ₂₇											-391
C≡O ₂₈											-428
C≡O ₂₉											-465
C≡O ₃₀											-502
C≡O ₃₁											-539
C≡O ₃₂											-576
C≡O ₃₃											-613
C≡O ₃₄											-650
C≡O ₃₅											-687
C≡O ₃₆											-724
C≡O ₃₇											-761
C≡O ₃₈											-798
C≡O ₃₉											-835
C≡O ₄₀											-872
C≡O ₄₁											-909
C≡O ₄₂											-946
C≡O ₄₃											-983
C≡O ₄₄											-1020
C≡O ₄₅											-1057
C≡O ₄₆											-1094
C≡O ₄₇											-1131
C≡O ₄₈											-1168
C≡O ₄₉											-1205
C≡O ₅₀											-1242
C≡O ₅₁											-1279
C≡O ₅₂											-1316
C≡O ₅₃											-1353
C≡O ₅₄											-1390
C≡O ₅₅											-1427
C≡O ₅₆											-1464
C≡O ₅₇											-1501
C≡O ₅₈											-1538
C≡O ₅₉											-1575
C≡O ₆₀											-1612
C≡O ₆₁											-1649
C≡O ₆₂											-1686
C≡O ₆₃											-1723
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C≡O ₉₁											-2759
C≡O ₉₂											-2796
C≡O ₉₃											-2833
C≡O ₉₄											-2870
C≡O ₉₅											-2907
C≡O ₉₆											-2944
C≡O ₉₇											-2981
C≡O ₉₈											-3018
C≡O ₉₉											-3055
C≡O ₁₀₀											-3092

1. Write the balanced equation for the combustion of 1 mole of glucose, C₆H₁₂O₆, completely in O₂.



2. Draw the structural formula showing Lewis structures for each of the molecules involved.

B = 1.07, C = 2.25, O = 3.44, H = 2.20, N = 3.04, F = 4.04, Cl = 3.16, Br = 2.96, I = 2.66, At = 2.20



3. Calculate the bond energies of the reactants.

Bond	kJ/mol	kJ/mol
C-H	413	3718
C-O	336	3718
O-O	498	1275

4. Calculate the bond energies of the products.

Bond	kJ/mol	kJ/mol
C=O	799	3794
O-H	463	2728

Total Bond Energy (Reactants)	-4627 kJ	Total Bond Energy (Products)	-5479 kJ
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5. Calculate the Heat Energy released. Bond Energy = _____

What is the Heat Energy released (combustion)? -1852 kJ

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Chemical Bonds Harry B. Gray, 1994-12-05 This profusely illustrated book by a world renowned chemist and award winning chemistry teacher provides science students with an introduction to atomic and molecular structure and bonding This is a reprint of a book first published by Benjamin Cummings 1973

Structure and Bonding Jack Barrett, 2001 This book explains in non mathematical terms where possible the factors that govern covalent bond formation the lengths and strengths of bonds and molecular shapes

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Bonding and Molecular Structure Russell S. Drago, 1971

The Shape and Structure of Molecules Charles Alfred Coulson, 1973

A Pictorial Approach to Molecular Bonding John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as NH_3 are somehow different from those which apply to aromatic systems such as C_6H_6 Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH_4 for example

Molecules and the Chemical Bond Henry A. Bent, 2011 MOLECULES AND THE CHEMICAL BOND Chemistry Simplified This highly original book by a famous chemistry teacher about general chemistry in a new key may change how teachers teach Atomic Theory The Mole Concept and Avogadro's Constant The Gas Laws Solving Problems in Chemical Stoichiometry The Saturation and Directional Character of Chemical Affinity The Pauli Exclusion Principle Linnett's Double Spin Set Theory Pauling's Rules of Crystal Chemistry The Octet Rule Lewis Structures for O_2 NO CO SO_2 and SO_3 Construction of Bond Diagrams VSEPR Theory Dative Bonding Multicenter Bonding Bonding in Metals pH Calculations The

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