

## Thermochemistry Review - Answers

1. Using the data below, calculate the amount of energy required to melt a 16.2 g aluminum pop can. Assume that the aluminum is at an initial temperature of 20°C.

- Molar heat of fusion of Al(s) = 10.7 kJ/mol
- Melting point of Al(s) = 660°C
- Heat capacity of Al(s) = 0.900 J/g°C

$$\text{mol} = 16.2 \text{ g} \div 26.98154 \text{ g/mol} = 0.600 \text{ mol}$$

$$\Delta H = (20^\circ\text{C} \rightarrow 660^\circ\text{C}) + (\text{melt})$$

$$\Delta H = (mc\Delta T) + (n\Delta H_{\text{fus}})$$

$$\Delta H = (16.2 \text{ g})(0.900 \text{ J/g}^\circ\text{C})(640^\circ\text{C}) + (0.600 \text{ mol})(10.7 \text{ kJ/mol})$$

$$\Delta H = 9331.2 \text{ J} + 6.42 \text{ kJ}$$

$$\Delta H = 9.33 \text{ kJ} + 6.42 \text{ kJ}$$

$$\Delta H = \mathbf{15.8 \text{ kJ}}$$

2. Given the thermochemical equations:



Calculate the  $\Delta H$  for the reaction:  $\text{CH}_3\text{OH}(\text{l}) + \frac{3}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$



# Ch Thermochemistry Answers

**Siddappa N.Byrareddy**



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Solutions Manual to Accompany Elements of Physical Chemistry C. A. Trapp, M. P. Cady, 2007 The Solutions manual to accompany Elements of Physical Chemistry 4e contains full worked solutions to all end of chapter exercises featured in the book

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## **Table of Contents Ch Thermochemistry Answers**

1. Understanding the eBook Ch Thermochemistry Answers
  - The Rise of Digital Reading Ch Thermochemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Ch Thermochemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ch Thermochemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ch Thermochemistry Answers
  - Personalized Recommendations

- Ch Thermochemistry Answers User Reviews and Ratings
- Ch Thermochemistry Answers and Bestseller Lists
- 5. Accessing Ch Thermochemistry Answers Free and Paid eBooks
  - Ch Thermochemistry Answers Public Domain eBooks
  - Ch Thermochemistry Answers eBook Subscription Services
  - Ch Thermochemistry Answers Budget-Friendly Options
- 6. Navigating Ch Thermochemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Ch Thermochemistry Answers Compatibility with Devices
  - Ch Thermochemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Ch Thermochemistry Answers
  - Highlighting and Note-Taking Ch Thermochemistry Answers
  - Interactive Elements Ch Thermochemistry Answers
- 8. Staying Engaged with Ch Thermochemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ch Thermochemistry Answers
- 9. Balancing eBooks and Physical Books Ch Thermochemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ch Thermochemistry Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Ch Thermochemistry Answers
  - Setting Reading Goals Ch Thermochemistry Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ch Thermochemistry Answers
  - Fact-Checking eBook Content of Ch Thermochemistry Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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