

**10****CHEMICAL QUANTITIES****Practice Problems**

*On a piece of paper, solve the following problems.*

***YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!***

**SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER**

1. What is the molar mass of sucrose ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ )?
2. What is the molar mass of each of the following compounds?
  - a. phosphorus pentachloride ( $\text{PCl}_5$ )
  - b. uranium hexafluoride ( $\text{UF}_6$ )
3. Calculate the molar mass of each of the following ionic compounds:
  - a.  $\text{KMnO}_4$
  - b.  $\text{Ca}_3(\text{PO}_4)_2$
4. How many moles is  $3.52 \times 10^{24}$  molecules of water?
5. How many atoms of zinc are in 0.60 mol of zinc?
6. What is the mass of 1.00 mol of oxygen ( $\text{O}_2$ )?

**SECTION 10.2 MOLE–MASS AND MOLE–VOLUME RELATIONSHIPS**

1. What is the molar mass of each of the following compounds?
  - a.  $\text{C}_6\text{H}_{12}\text{O}_6$
  - b.  $\text{NaHCO}_3$
  - c.  $\text{C}_3\text{H}_8$
  - d.  $\text{KNH}_4\text{SO}_4$
2. Calculate the mass in grams of each of the following:
  - a. 8.0 mol lead oxide ( $\text{PbO}$ )
  - b. 0.75 mol hydrogen sulfide ( $\text{H}_2\text{S}$ )
  - c. 0.00100 mol silicon tetrahydride ( $\text{SiH}_4$ )
  - d.  $1.50 \times 10^{-2}$  mol molecular oxygen ( $\text{O}_2$ )
  - e. 2.30 mol ethylene glycol ( $\text{C}_2\text{H}_6\text{O}_2$ )
3. How many grams are in 1.73 mol of dinitrogen pentoxide ( $\text{N}_2\text{O}_5$ )?
4. How many grams are in 0.658 mol of calcium phosphate [ $\text{Ca}_3(\text{PO}_4)_2$ ]?
5. Calculate the number of moles in each of the following:
  - a. 0.50 g sodium bromide ( $\text{NaBr}$ )
  - b. 13.5 g magnesium nitrate [ $\text{Mg}(\text{NO}_3)_2$ ]
  - c. 1.02 g magnesium chloride ( $\text{MgCl}_2$ )
  - d. 0.00100 g monochloromethane ( $\text{CH}_3\text{Cl}$ )
  - e.  $1.50 \times 10^{-3}$  g propylene glycol [ $\text{C}_3\text{H}_8(\text{OH})_2$ ]
6. A chemist plans to use 435.0 grams of ammonium nitrate ( $\text{NH}_4\text{NO}_3$ ) in a reaction. How many moles of the compound is this?
7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine ( $\text{C}_{20}\text{H}_{24}\text{N}_2\text{O}_2$ ). What mass, in grams, should the laboratory technician obtain in order to make the solution?
8. What is the volume at STP of 2.66 mol of methane ( $\text{CH}_4$ ) gas?
9. How many moles is 135 L of ammonia ( $\text{NH}_3$ ) gas at STP?

# Chapter 10 Chemical Quantities Answers

**Dennis D. Staley, Antony C. Wilbraham**



## Chapter 10 Chemical Quantities Answers:

Basic Chemistry Steven S Zumdahl,1996      *Chemistry: Core Concepts, 3rd Edition* Allan Blackman,Daniel Southam,Gwendolyn Lawrie,Natalie Williamson,Christopher Thompson,2024-01-09 The third edition of Chemistry Core Concepts Blackman et al has been developed by a group of leading chemistry educators for students entering university with little or no background in chemistry Available as a full colour printed textbook with an interactive eBook code this title enables every student to master concepts and succeed in assessment Lecturers are supported with an extensive and easy to use teaching and learning package      **Chemistry** Hall,1995-12      **Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual** Leo J. Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

**Introductory Chemistry Problems: The Chemistry And Physics Behind The Numbers** Michael E Green,2025-07-18 The essential point of the book is to show students in Introductory Chemistry classes how to obtain answers that are at least reasonable and to realize when there must have been an error because the answer falls outside the range of possible values Students often leave ridiculous answers because they simply plug into equations they have memorized correctly or incorrectly They have not learned to think about the meaning of the numbers that appear in their calculations This book attempts to show students how to correct this      *Study Guide/Selected Solutions Manual* Julie R. Frentrup,2002-06 Contains a brief overview of every chapter review of skills self tests and the answers and detailed solutions to all odd numbered end of chapter problems in the text book      *Essentials of General, Organic, and Biochemistry* Denise Guinn,2019-03-05 Guinn s Essentials of General Organic and Biochemistry uses health and medicine as the framework for learning the fundamentals of chemistry in this student centered one semester text The newly revised 3rd edition focuses on core concepts and necessary math skills and features a revamped organization to align with traditional course organization and shorter more condensed chapters Easily digestible content and medical applications help reduce student anxiety and make chemistry meaningful for students preparing for future careers in nursing and other allied health professions Paired with SaplingPlus and an

embedded eBook students will be able to focus their study with adaptive quizzing and understand the relevance of chemistry through videos animations and case studies      **Essentials of General, Organic, and Biochemistry Digital Update**

Denise Guinn,2024-01-17 Guinn s Essentials of General Organic and Biochemistry 3e Digital Update is the best intro available for teaching the basics of chemistry to students preparing for careers in nursing and other allied health professions

**Basic Chemical Principles** Donada Peters,Edward I. Peters,1988      **Study Guide for Introductory Chemistry** Iris Stovall,1990      Chemistry John S. Phillips,Cheryl Wistrom,2000      *Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry* Iris Stovall,Steven S. Zumdahl,1996      Chemistry for the Biosciences Jonathan Crowe,Tony Bradshaw,Paul M. S. Monk,2006 Leading students through the essential concepts that are central to understanding biological systems this text uses everyday examples and analogies to build their confidence in an often daunting subject By focusing on the key themes that unify the subject it shows how integral chemistry is to the biosciences

*Chemistry* Zumdahl,Steven S. Zumdahl,2003-07 Contains chapter discussions key term definitions and practice chapter tests      **Introductory Chemistry** Steven S. Zumdahl,1993      Process Operations Safety Darryl M. Yoes,2025-02-05 Provides crucial lessons in process safety operations drawing from 100 global case studies Written from an operator s perspective Process Operations Safety provides valuable information and education on the fundamentals of process operations safety by providing background on process safety and key leading operational management and equipment failures that have led to catastrophic process safety incidents including loss of life Written by an expert with more than five decades of industry experience this book enables readers to learn how simple jobs that they perform every day can lead to catastrophic safety incidents without proper caution protocol and attention A self learning quiz is provided near each chapter s end with answers to all questions provided in the Appendix A listing of additional resources or reference material many with internet links is also included at the end of each chapter Readers will find Principles of process safety properties of hydrocarbons vapor cloud explosions VCE and boiling liquid expanding vapor explosions BLEVE Most frequent causes of significant process safety events in refining and petrochemical industries Causal factors in over 100 global case studies of operations and incidents divided into thirty five subchapters with several examples for each explaining what happened and what could have happened Key lessons learned written in simple terms using descriptions without jargon or complicated formulas Process Operations Safety is an essential learning resource for petroleum refining and petrochemical plant operators line supervisors and critical support staff with field responsibility such as process and mechanical engineers along with advanced students at community and four year colleges and technical trade schools taking a process operations course

**General Organic and Biological Chemistry** Kenneth W. Raymond,2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to

have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features     Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall,Donald J. DeCoste,2000     **Basic Concepts of Chemistry** Leo J. Malone,Theodore O. Dolter,2012-04-13 Basic Concepts of Chemistry was originally written over thirty years ago to address the needs of general chemistry students with little or no background in chemistry Over time the text has evolved beyond purposes solely aimed at a preparatory chemistry course For some preparatory chemistry students a main sequence in general chemistry may follow but for others a semester of organic and biochemistry may follow Other students enroll to simply satisfy a basic science or chemistry requirement The text was written with a level and functionality designed to accommodate the needs of each of these varied groups of students Basic Concepts of Chemistry was designed with a flexibility that allows instructors to emphasize or omit certain clearly delineated sections The mission of the text has evolved in response to the increased diversity of students and the emphasis on outcomes assessment     *Essentials of Chemistry* Dennis D. Staley,Anthony C. Wilbraham,1984

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