Study Guide Chemical Reactions

- Define oxidation number and provide an example.
- Define binary compound and provide an example.
- 3. Discuss which types of elements lose electrons and which gain electrons.
- 4. What are oxidation numbers and how can you use the periodic table to determine them?
- Name four elements that tend to have more than one oxidation number.
- 6. Discuss the charges of ions standing alone and when they are combined.
- 7. What do you do if you have two ions in an equation with different exidation numbers?
- 8. What is a coefficient and what does it mean? Provide an example.
- 10. Name this binary ionic compound: N₂O
 11. What is meant by oxidation?

9. Balance the following equation:

- 12. What is meant by reduction?
- 13. Describe in detail a redox reaction.
- 14. Identify which element was oxidized and which was reduced in the equation you balanced above.

 $H_2O + CO_2 \rightarrow C_6H_{12}O_6 + O_2$

- 15. Distinguish between reactants and products.
- 16. What are the four steps to balancing an equation?
- Describe exothermic and endothermic reactions.
- Draw an exothermic graph and an endothermic graph.
- 19. What is meant by the activation energy of a reaction?
- 20. How does a catalyst affect a reaction?
- 21. Is the catalyst used up?
- 22. What is a chemical reaction?
- 23. Name five indicators that a chemical reaction has occurred.
- 24. Who was Antoine Lavoisier and what was his contribution to chemistry?
- 25. What is a chemical formula?
- 26. What does a chemical formula tell you?
- 27. Discuss chemical names.
- 28. Write the chemical formula for each of the following:

(a) four hydrogens and one carbon (methane); (b) carbon dioxide; (c)dihydrogen oxide

- List all of the noble gases and discuss their unique characteristics.
- 30. Why do atoms tend to bond with each other?
- 31. What is a chemical bond?
- 32. What is an ion?
- 33. How do ions form?
- Describe ionic bonding using NaCl as your example.
- 35. What is a covalent bond?
- 36. What is a polar covalent bond?
- 37. What is the advantage of a molecule being polar? Use water as an example.
- 38. What gives water its ability to dissolve substances?
- 39. How do multiple bonds form? Use N₂ as your example and discuss the movement of electrons.
- 40. What types of elements tend to bond with each other? Use the periodic table.
- 41. What is metallic bonding and how does it work?
- 42. How do ions form?

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Daniela Niemeyer

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