

13

STATES OF MATTER

SECTION 13.1 THE NATURE OF GASES (pages 385–389)

This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance.

► Kinetic Theory and a Model for Gases (pages 385–386)

1. The energy an object has because of its motion is called kinetic energy.
2. Circle the letter of each sentence that is true about the assumptions of the kinetic theory concerning gases.
 - a. A gas is composed of particles with insignificant volume that are relatively far apart from each other.
 - b. Strong attractive forces exist between particles of a gas.
 - c. Gases tend to collect near the bottom of a container.
 - d. The paths of uninterrupted travel of particles in a gas are relatively short because the particles are constantly colliding with each other or with other objects.
3. Is the following statement true or false? According to the kinetic theory, collisions between particles in a gas are perfectly elastic because kinetic energy is transferred without loss from one particle to another, and the total kinetic energy remains constant. true

► Gas Pressure (pages 386–387)

4. Gas pressure results from the force exerted by a gas per unit surface area of an object.
5. Simultaneous collisions of billions of particles in a gas with an object result in gas pressure.
6. What force holds the particles of air in Earth's atmosphere? gravity
7. What kind of pressure is measured with a barometer? atmospheric pressure

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D Kirk



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Table of Contents Chapter 10 Chemistry Study Guide Answers Prentice Hall

1. Understanding the eBook Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - The Rise of Digital Reading Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Personalized Recommendations
 - Chapter 10 Chemistry Study Guide Answers Prentice Hall User Reviews and Ratings
 - Chapter 10 Chemistry Study Guide Answers Prentice Hall and Bestseller Lists

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 - Chapter 10 Chemistry Study Guide Answers Prentice Hall eBook Subscription Services
 - Chapter 10 Chemistry Study Guide Answers Prentice Hall Budget-Friendly Options
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7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Highlighting and Note-Taking Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Interactive Elements Chapter 10 Chemistry Study Guide Answers Prentice Hall
8. Staying Engaged with Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 10 Chemistry Study Guide Answers Prentice Hall
9. Balancing eBooks and Physical Books Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 10 Chemistry Study Guide Answers Prentice Hall
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Setting Reading Goals Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Fact-Checking eBook Content of Chapter 10 Chemistry Study Guide Answers Prentice Hall
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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