

Matter and Change, pp. 1-4

- 2.
- 3.
- 4.
9. mass and volume
10. group or family
11. homogeneous
12. products
13. metal
14. solid state
15. chemical properties
16. atom
17. compound
18. nonmetal
19. mixture
20. pure substance
21. basic
22. metalloid
23. reactants: carbon and oxygen; product: carbon dioxide
24. reactant: mercury(II) oxide; products: mercury and oxygen
25. f
26. g
27. c
28. a
29. b
30. i
31. d
32. h
33. e
34. physical change
35. chemical change
36. physical change
37. physical change
38. chemical change
39. physical change
40. In a solid, particles are packed together in relatively fixed positions. The particles vibrate about a fixed point. In a liquid, particles are close together but can flow around one another. In a gas, particles are at great distances from one another, compared to the particles of liquids and solids.
41. The composition of a pure substance is the same throughout and does not vary from sample to sample. Pure substances include elements and compounds. A mixture contains more than one substance and can vary in composition from one sample to another and sometimes within different parts of the same sample.
42. Although energy can be absorbed or released in a

5. c
5. c
7. c
3. a

change, it is not destroyed or created. It simply takes a different form.

43. Homogeneous mixtures are uniform in composition; heterogeneous mixtures are not.
44. An element cannot be broken down, but water can be decomposed into hydrogen and oxygen by passing an electrical current through it (electrolysis).

Measurements and Calculations, pp. 5-8

- a
5. c
2. c
6. c
3. c
7. a
4. a
8. a
9. time
10. mass
11. density
12. energy
13. length or distance
14. volume
15. area
16. qualitative
17. quantitative
18. qualitative
19. quantitative
20. 3.00×10^3 km/s
21. three
22. 0.026 g or 2.6×10^{-2} g
23. 2.5×10^{-4} L
24. quantity
25. derived units
26. conversion factor
27. 0.0432 kg or 4.32×10^{-2} kg
28. 5400 mL or 5.4×10^3 mL
29. 300. K
30. 1.05×10^3 J
31. 3.51×10^3 cm
32. A hypothesis is a testable statement that can be used to make predictions and to carry out further experiments. A theory is a broad generalization that explains a body of facts or phenomena.
33. Mass is a measure of the quantity of matter. Weight is a measure of the gravitational pull on matter. Mass does not depend on gravitational attraction.
34. Models are visual, verbal, or mathematical representations. They are used to explain how phenomena occur or how data or events are related.
35. The graph of two quantities that are directly proportional is a straight line through the origin.

Chapter 11 Chemistry Matter And Change Answer Key

Philip D. Rufe



Chapter 11 Chemistry Matter And Change Answer Key:

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Glencoe Chemistry Solving Problems A Chemistry Handbook Matter and Change Chemistry McGraw-Hill Staff, 2001-03
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general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features

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A dictionary of chemistry. (Second, Third suppl.). Henry Watts,1872 **Books in Print Supplement** ,2002

Modern Techniques in Computational Chemistry Enrico Clementi,1989 In a way the MOTECC 89 project started in the early sixties at the IBM Research Laboratory in San Jose California The six years of post doctoral research first with Giulio Natta on conductive polymers with Michael Kasha on spin orbit effects with Kenneth S Pitzer on high temperature molecules and thermo dynamics and with R S Mulliken in the quantum chemistry of small molecules had demonstrated pragmatically the importance of a broad based research and also let me taste some of the excitement to be derived from interdisciplinarity Thus when I started to gather a department in the newly opened IBM Research Laboratory in San Jose California I purposely named it Large Scale Scientific Computation Department avoiding reference to chemistry physics statistical mechanics or fluid dynamics which were our main tasks In the sixties interdisciplinarity was more and more recognized as a most important if not necessary avenue to cope with the technical needs of our society However at that time interdisciplinarity was synonymous with team work and true interdisciplinarity was a formidably difficult objective Although I headed an excellent group of scientists with different backgrounds and there was much progress and creativity still each one of us was more or less conducting his own research in his own field with occasional cross field partnerships and with some of the computational techniques as our common base Later in 1974 I made a second attempt this time starting a new department at the Donegani Institute Montedison in Novara Italy

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