

11.1 - Describing Chemical Reactions

1. What two substances reacted in the Hindenburg crash?
2. In general, what happens in a **chemical reaction**?
3. Look at Figure 11.1. What are some of the evidences of **chemical change** (a.k.a. chemical reaction)?
4. Give an example of a **word equation** for a chemical reaction (you can make one up or choose one from the book).
5. How is a **chemical equation** different from a word equation?
6. What is missing in a skeleton equation?
7. What symbols are used to represent the phases of substances in chemical reactions?
8. What does the symbol (aq) mean? _____
9. Use the space below to do Practice Problems #1 and #2.
1) _____ 2) _____
10. A complete description of a reaction must include _____,
but also _____.
11. Define the following:
 - a. **coefficients** -
 - b. **subscripts** -
 - c. **balanced equation** -
12. How does the **law of conservation of mass** relate to balancing reaction equations?
13. In a chemical reaction, what **MUST** be the same for the reactants and the products?
(There are two great answers to this question!)

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