

- (a) Decomposition of ferrous sulphate
- (b) Dilution of sulphuric acid
- (c) Dissolution of sodium hydroxide in water
- (d) Dissolution of ammonium chloride in water

5. Identify the reducing agent in the following reactions

- (a)  $4\text{NH}_3 + 5\text{O}_2 \rightarrow 4\text{NO} + 6\text{H}_2\text{O}$
- (b)  $\text{H}_2\text{O} + \text{F}_2 \rightarrow \text{HF} + \text{HOF}$
- (c)  $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
- (d)  $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

6. Identify the oxidising agent (oxidant) in the following reactions

- (a)  $\text{Pb}_3\text{O}_4 + 8\text{HCl} \rightarrow 3\text{PbCl}_2 + \text{Cl}_2 + 4\text{H}_2\text{O}$
- (b)  $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
- (c)  $\text{CuSO}_4 + \text{Zn} \rightarrow \text{Cu} + \text{ZnSO}_4$
- (d)  $\text{V}_2\text{O}_5 + 5\text{Ca} \rightarrow 2\text{V} + 5\text{CaO}$
- (e)  $3\text{Fe} + 4\text{H}_2\text{O} \rightarrow \text{Fe}_3\text{O}_4 + 4\text{H}_2$
- (f)  $\text{CuO} + \text{H}_2 \rightarrow \text{Cu} + \text{H}_2\text{O}$

7. Write the balanced chemical equations for the following reactions

- (a) Sodium carbonate on reaction with hydrochloric acid in equal molar concentrations gives sodium chloride and sodium hydrogencarbonate.
- (b) Sodium hydrogencarbonate on reaction with hydrochloric acid gives sodium chloride, water and liberates carbon dioxide.
- (c) Copper sulphate on treatment with potassium iodide precipitates cuprous iodide ( $\text{Cu}_2\text{I}_2$ ), liberates iodine gas and also forms potassium sulphate.

8. A solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance is formed. Write the chemical reaction involved and also mention the type of

# Chapter 10 1 Reactions And Equations Worksheet Answers

**SJ Ball**



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## **Table of Contents Chapter 10 1 Reactions And Equations Worksheet Answers**

1. Understanding the eBook Chapter 10 1 Reactions And Equations Worksheet Answers
  - The Rise of Digital Reading Chapter 10 1 Reactions And Equations Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 10 1 Reactions And Equations Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 10 1 Reactions And Equations Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 10 1 Reactions And Equations Worksheet Answers
  - Personalized Recommendations
  - Chapter 10 1 Reactions And Equations Worksheet Answers User Reviews and Ratings
  - Chapter 10 1 Reactions And Equations Worksheet Answers and Bestseller Lists

5. Accessing Chapter 10 1 Reactions And Equations Worksheet Answers Free and Paid eBooks
  - Chapter 10 1 Reactions And Equations Worksheet Answers Public Domain eBooks
  - Chapter 10 1 Reactions And Equations Worksheet Answers eBook Subscription Services
  - Chapter 10 1 Reactions And Equations Worksheet Answers Budget-Friendly Options
6. Navigating Chapter 10 1 Reactions And Equations Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 10 1 Reactions And Equations Worksheet Answers Compatibility with Devices
  - Chapter 10 1 Reactions And Equations Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 10 1 Reactions And Equations Worksheet Answers
  - Highlighting and Note-Taking Chapter 10 1 Reactions And Equations Worksheet Answers
  - Interactive Elements Chapter 10 1 Reactions And Equations Worksheet Answers
8. Staying Engaged with Chapter 10 1 Reactions And Equations Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 10 1 Reactions And Equations Worksheet Answers
9. Balancing eBooks and Physical Books Chapter 10 1 Reactions And Equations Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 10 1 Reactions And Equations Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 10 1 Reactions And Equations Worksheet Answers
  - Setting Reading Goals Chapter 10 1 Reactions And Equations Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 10 1 Reactions And Equations Worksheet Answers
  - Fact-Checking eBook Content of Chapter 10 1 Reactions And Equations Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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