

Guide to Chapter 10. Liquids and Solids

Answers in red

We will spend three lecture days on this chapter.

Read the introductory paragraph to Chapter 10.

Read Section 10.1 Polar Covalent Bonds and Dipole Moments.

Learning Objective 1: Given the formula of a molecular compound determine if it is polar or non polar covalent.

Learning Objective 2: Know the definition of STP. Know the standard molar volume (22.4 L)

Do Problems 1 - 4 at the end of the section.

Do the following end-of-chapter problems: 22, 30, 36, 38

Read Section 10.2 Intermolecular Forces and Section 10.3 Some Properties of Liquids

Learning Objective 3: Given the formula of a compound, determine the major attractive force. (ionic = ion-ion and ion-dipole) and (molecular and covalent = dipole-dipole, hydrogen bonding, or London dispersion forces).

Learning Objective 4: Given a list of compound formulas rank their attractive forces in order of strength.

Learning Objective 5: Describe the relationship between size and strength of London dispersion forces.

Learning Objective 6: Describe the relationship between size and strength of London dispersion forces.

Do Problems 5 and 6 at the end of Section 10.2.

Do the following end-of-chapter problems: 32, 34, 40

Problem Club Question A. (a) Create a graph (general appearance only) with *MM* on the x-axis and boiling points on the y-axis for these similar compounds: NH_3 , PH_3 , AsH_3 , and SbH_3 . (b) Repeat for CH_4 , SiH_4 , GeH_4 , and SnH_4 .

Answer: boiling points (°C): (a) $\text{NH}_3 = -33$, $\text{PH}_3 = -87$, $\text{AsH}_3 = -55$, $\text{SbH}_3 = -17$. (b) $\text{CH}_4 = -164$, $\text{SiH}_4 = -111$, $\text{GeH}_4 = -89$, and $\text{SnH}_4 = -52$

Problem Club Question B. Why is iodine a solid?

*Answer: very substantial London-dispersion forces due to the large *MM**

Problem Club Question C. Circle the chemical with the highest melting point:

Answers in red

(a) F_2 or Cl_2

(b) F_2 or PCl_3

(c) Cl_2 or PCl_3

(d) NaCl or Cl_2

(e) HF or HBr

(f) $\text{CH}_3\text{CH}_2\text{CH}_3$ or $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$

Problem Club Question D. True or False:

Answers in red

T F

Covalent bonds are weaker than London-dispersion forces

T F

Dipole forces are stronger than London-dispersion forces

T F

Only non-polar molecules have London-dispersion forces

T F

CaS and KCl have similar *MMs*, however CaS should have a much larger mp.

T F

The heat of vaporization increases with *MM* for the series $(\text{BH}_3)_n$

Chapter 10 Liquids And Solids Answers

Patrick Vollmar



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Solutions Guide for Chemistry, Third Edition, by Steven S. Zumdahl Kenneth C. Brooks, 1993

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