

# Chemical Reactions

## Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- |               |   |
|---------------|---|
| <u>yes</u>    | 1. A tomato smells rotten.  |
| <u>no</u>     | 2. A drinking glass breaks into smaller pieces.                     |
| <u>no</u>     | 3. A piece of ice melts.  |
| <u>yes</u>    | 4. Drain cleaner is mixed with water and the solution becomes warm. |
| <u>yes</u>    | 5. Candle wax burns.  |
| <u>yes/no</u> | 6. Molten candle wax solidifies.                                    |
| <u>yes</u>    | 7. Green leaves turn yellow and red as the seasons change.          |
| <u>yes</u>    | 8. Baking powder produces a gas that makes a cake rise.             |

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** reactant, oxygen is a(n) **(10)** reactant, and water vapor is a(n) **(11)** product. In a chemical equation for this reaction, a(n) **(12)** arrow is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** plus sign is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** (g) symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** (g) symbol is used for the state of oxygen, and a(n) **(16)** (g) symbol is used for the state of water vapor.

# Concept Review Describing Chemical Reactions Answer Key

**Kathleen Armour**



## Concept Review Describing Chemical Reactions Answer Key:

**Chemistry<sup>3</sup>** Andrew Burrows, John Holman, Simon Lancaster, Andrew Parsons, Tina Overton, Gwen Pilling, Gareth Price, 2021 Chemistry is widely considered to be the central science it encompasses concepts on which all other branches of science are developed Yet for many students entering university gaining a firm grounding in chemistry is a real challenge Chemistry<sup>3</sup> responds to this challenge providing students with a full understanding of the fundamental principles of chemistry on which to build later studies Uniquely amongst the introductory chemistry texts currently available Chemistry<sup>3</sup> s author team brings together experts in each of organic inorganic and physical chemistry with specialists in chemistry education to provide balanced coverage of the fundamentals of chemistry in a way that students both enjoy and understand The result is a text that builds on what students know already from school and tackles their misunderstandings and misconceptions thereby providing a seamless transition from school to undergraduate study Written with unrivalled clarity students are encouraged to engage with the text and appreciate the central role that chemistry plays in our lives through the unique use of real world context and photographs Chemistry<sup>3</sup> tackles head on two issues pervading chemistry education students mathematical skills and their ability to see the subject as a single unified discipline Instead of avoiding the maths Chemistry<sup>3</sup> provides structured support in the form of careful explanations reminders of key mathematical concepts step by step calculations in worked examples and a Maths Toolkit to help students get to grips with the essential mathematical element of chemistry Frequent cross references highlight the connections between each strand of chemistry and explain the relationship between the topics so students can develop an understanding of the subject as a whole Digital formats and resources Chemistry<sup>3</sup> is available for students and institutions to purchase in a variety of formats and is supported by online resources The e book offers a mobile experience and convenient access along with functionality tools navigation features and links that offer extra learning support [www.oxfordtextbooks.co.uk/ebooks](http://www.oxfordtextbooks.co.uk/ebooks) The e book also features interactive animations of molecular structures screencasts in which authors talk step by step through selected examples and key reaction mechanisms and self assessment activities for each chapter The accompanying online resources will also include for students DT Chapter 1 as an open access PDF DT Chapter summaries and key equations to download to support revision DT Worked solutions to the questions in the book The following online resources are also provided for lecturers DT Test bank of ready made assessments for each chapter with which to test your students DT Problem solving workshop activities for each chapter for you to use in class DT Case studies showing how instructors are successfully using Chemistry<sup>3</sup> in digital learning environments and to support innovative teaching practices DT Figures and tables from the book

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*General Chemistry* Stephen D. Bresnick,1996 A study guide in question and answer format for basic chemistry *Holt Chemistry* Ralph Thomas Myers,2004 **Chemical Interactions** ,2005 *Merrill Chemistry* Robert C. Smoot,Smoot,Richard G. Smith,Jack Price,1998 *Contemporary Chemical Analysis* Judith F. Robinson,Kenneth A. Robinson,1998 This book provides a readable yet rigorous introduction to analytical methods with a focus on problem solving skills It stresses the fundamental concepts of chemical analysis and through examples from current journals and other science media shows how the principles and practice of analytical chemistry are used to produce answers to questions in all areas of scientific study and practice Features a balance of topics that is closer to contemporary analytical practice than those covered by other books Introduces the tools that are ubiquitous in analytical chemistry e g statistics sampling and sample preparation Discusses methods depending on chemical kinetics which are so widely used in medicine and biology Features a number of problems that call for the use of a spreadsheet to generate data which is then plotted to show trends Includes answers for all numerical problems in an appendix **Foundations of College Chemistry** Morris Hein,Susan Arena,1993 This is the Student Study Guide to accompany Hein s Foundations of College Chemistry Alternate Edition 12th Edition **Industrial Engineering** George Worthington,1900 **Concepts in Biology** Eldon Enger,Frederick Ross,David Bailey,2008-01-16 Enger Ross Bailey Concepts in Biology is a relatively brief introductory general biology text written for students with no previous science background The authors strive to use the most accessible vocabulary and writing style possible while still maintaining scientific accuracy The text covers all the main areas of study in biology from cells through ecosystems Evolution and ecology coverage are combined in Part Four to emphasize the relationship between these two main subject areas The new 13th edition is the latest and most exciting revision of a respected introductory biology text written by authors who know how to reach students through engaging writing interesting issues and applications and accessible level Instructors will appreciate the book s scientific accuracy complete coverage and extensive supplement package

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