Conductivity of Solutions: The Effect of Concentration

If an ionic compound is dissolved in water, it dissociates into ions and the resulting solution will conduct electricity. Dissolving solid sodium chloride in water releases ions according to the equation:

$$NaCl(s) \longrightarrow Na'(aq) + Cl'(aq)$$

In this experiment, you will study the effect of increasing the concentration of an ionic compound on conductivity. Conductivity will be measured as concentration of the solution is gradually increased by the addition of concentrated NaCl drops. The same procedure will be used to investigate the effect of adding solutions with the same concentration (1.0 M), but different numbers of ions in their formulas: aluminum chloride, AlCl₃, and calcium chloride, CaCl₂. A computer-interfaced Conductivity Probe will be used to measure conductivity of the solution. Conductivity is measured in microsiemens (µS).

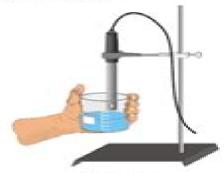


Figure 1

MATERIALS

Power Macintosh or Windows PC Vernier computer interface Logger Pro Vernier Conductivity Probe ring stand utility clamp 100-mL beaker distilled water 1.0 M NaCl solution 1.0 M CaCl₂ solution 1.0 M AlCl₃ solution stirring rod

PROCEDURE

- 1. Obtain and wear goggles.
- Prepare the computer for data collection by opening the file in the Experiment 14 folder of Chemistry with Computers. The vertical axis will have conductivity scaled from 0 to 2000 µS. The horizontal axis will have volume scaled from 0 to 8 drops.

Conductivity Lab Ap Chemistry

Lei Huang

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