

Chapter 20 Electrode potentials and electrochemical cells

Practice questions

Question number	Answer	Marks	Guidance
1 (a)	By definition	1	Allow 'set to this value'
1 (b)	1.23V	1	Allow + or -
1 (c)	$\text{Pt} \text{H}_2(\text{g}) \text{OH}^-(\text{aq}), \text{H}_2\text{O}(\text{l}) \text{O}_2(\text{g}) \text{H}_2\text{O}(\text{l}), \text{OH}^-(\text{aq}) \text{Pt}$ Correct but with Pt missing Includes Pt with correct representation	1 1	H_2O not essential, allow reverse order
1 (d)	Uses $\text{O}_2 + 2\text{H}_2\text{O} + 4\text{e}^- \rightarrow 4\text{OH}^-$ And (2x) $2\text{OH}^- + \text{H}_2 \rightarrow 2\text{H}_2\text{O} + 2\text{e}^-$ $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$	1 1	
1 (e)	Increases the surface area (so reaction faster)	1	
1 (f)	Overall reaction is the same ($2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$)	1	Or shows e.m.f. is the same
1 (g)	Hydrogen and oxygen supplied continuously OR Can be operated without stopping to recharge	1	Or can be refuelled quickly Allow any one mark
1 (h)	Hydrogen may need to be made using an energy source that is not 'carbon neutral'	1	
2 (a)	Zn	1	This is always the species with the most negative E value (it is oxidised itself).
2 (b) (i)	Fe^{2+}	1	
2 (b) (ii)	Cl_2	1	
2 (c) (i)	The standard electrode potential: 1.25V	1	Remember that: $E_{\text{cell}} = E_{\text{right}} - E_{\text{left}}$
2 (c) (ii)	$\text{Ti}^{3+} + 2\text{Fe}^{2+} \rightarrow 2\text{Fe}^{3+} + \text{Ti}^+$	2	One mark for a balanced equation and one for the correct direction.
3 (a)	$\text{Pt} \text{H}_2 \text{H}^+ \text{Fe}^{2+} \text{Fe}$ Note, allow one mark only for correct symbol in reverse: $\text{Fe} \text{Fe}^{2+} \text{H}^+ \text{H}_2 \text{Pt}$	2	Allow 1 for correct order of symbols but lose 2nd mark for a wrong phase boundary(s) / Pt missing / extra Pt on RHS, additional phase boundary Allow dashed lines for salt bridge Ignore state symbols Ignore 2 if used before H^+
3 (b)	Electron donor	1	Allow (species that) loses electrons Do not allow reference to electron pairs
3 (c)	Cl_2 / chlorine (Species on RHS / electron donor) has most positive / largest E° / has highest potential	1 1	If M1 blank or incorrect cannot score M2 Do not allow reference to e.m.f. or $E(\text{cell})$
3 (d) (i)	Cl / chlorine	1	
3 (d) (ii)	Chlorine +1 to chlorine 0	1	CE if chlorine not identified in 3(d)(i) Allow chlorine +1 to chlorine -1 (in Cl^-) Allow oxidation state decreases by one / two Allow oxidation state changes by -1 OR -2
3 (e)	$4\text{HOCl} + 4\text{H}^+ + 4\text{OH}^- \rightarrow 2\text{Cl}_2 + \text{O}_2 + 6\text{H}_2\text{O}$ OR $4\text{HOCl} \rightarrow 2\text{Cl}_2 + \text{O}_2 + 2\text{H}_2\text{O}$	2	Allow one mark for any incorrect equation that shows $\text{HOCl} \rightarrow \text{Cl}_2 + \text{O}_2$ Allow multiples Ignore state symbols Penalise one mark for uncanceled / uncombined species (eg $\text{H}_2\text{O} + \text{H}_2\text{O}$ instead of $2\text{H}_2\text{O}$)
3 (f)	e.m.f. = $0.40 - (-1.25) = 1.65\text{V}$ / $+1.65\text{V}$	1	Allow -1.65V

Ch20 Chemistry Study Guide Answers

Steven S. Zumdahl, Paul B. Kelter



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drug roughly 300 mg For this reason as well as the spectacular molecular architecture of Taxol synthetic organic chemists fiercely undertook efforts to synthesize it Five total syntheses of Taxol have thus far been reported Now a combination of isolation of a related metabolite from European yew needles and synthesis of Taxol from that intermediate supply the clinical demand This case clearly demonstrates the importance of synthesis and the use of organic chemistry It s just one of the many examples used in the text that will spark the interest of students and get them involved in the study of organic chemistry

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chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application Solutions Manual and Study Guide to Accompany Introduction to Organic Chemistry, 4th Ed Paul A. Bartlett, Judith G. Koch, 1992

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