

- _____ 13. If the spin of one electron in an orbital is clockwise, what is the spin of the other electron in that orbital?
 a. zero c. counterclockwise
 b. clockwise d. both clockwise and counterclockwise
- _____ 14. What types of atomic orbitals are in the third principal energy level?
 a. *s* and *p* only c. *s*, *p*, and *d* only
 b. *p* and *d* only d. *s*, *p*, *d*, and *f*
- _____ 15. What is the next atomic orbital in the series 1*s*, 2*s*, 2*p*, 3*s*, 3*p*?
 a. 2*d* c. 3*f*
 b. 3*d* d. 4*s*
- _____ 16. What is the number of electrons in the outermost energy level of an oxygen atom?
 a. 2 b. 4 c. 6 d. 8
- _____ 17. What is the electron configuration of potassium?
 a. $1s^2 2s^2 2p^2 3s^2 3p^2 4s^1$ c. $1s^2 2s^2 3s^2 3p^6 3d^1$
 b. $1s^2 2s^2 2p^{10} 3s^2 3p^3$ d. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^1$
- _____ 18. How many unpaired electrons are in a sulfur atom (atomic number 16)?
 a. 0 b. 1 c. 2 d. 3
- _____ 19. Which is the correct electron dot diagram for the element Silicon (Si)?
 A. $\cdot \ddot{\text{Si}} \cdot$ B. $:\ddot{\text{Si}}:$ C. $\text{Si} \cdot$ D. $:\text{Si}:$
 a. c.
 b. d.
- _____ 20. How many valence electrons are present in silicon?
 a. Two b. Four c. Six d. Eight

Completion: Complete each statement.

21. The _____ principle states that a maximum of two electrons may occupy a single atomic orbital, but only if the electrons have opposite spins.
22. The arrangement of electrons in an atom is called the _____ of an atom.
23. The element _____ has the following electron configuration: $[\text{Rn}]5f^6 6d^1 7s^2$.
24. The *f*-block elements are also referred to as the lanthanides and _____.
25. Xenon has _____ valence electrons.
26. Aluminum has _____ valence electrons.

Ch 5 Electrons In Atoms Test Answers

Stanley Chernicoff



Ch 5 Electrons In Atoms Test Answers:

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