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CHAPTER 2

ATOMS, MOLECULES, AND IONS

Problem Categories

Biological: 2.79, 2.80.

Conceptual: 2.31, 2.32, 2.33, 2.34, 2.61, 2.65, 2.71, 2.72, 2.81, 2.82, 2.93, 2.103, 2.113.

Descriptive: 2.24, 2.25, 2.26, 2.49, 2.50, 2.66, 2.70, 2.76, 2.84, 2.85, 2.86, 2.87, 2.88, 2.90, 2.91, 2.92, 2.94, 2.98, 2.101, 2.114.

Environmental: 2.122.

Organic: 2.47, 2.48, 2.69, 2.105, 2.107, 2.108, 2.115.

Difficulty Level

Easy: 2.7, 2.8, 2.13, 2.14, 2.15, 2.16, 2.23, 2.31, 2.32, 2.33, 2.43, 2.44, 2.45, 2.46, 2.47, 2.48, 2.62, 2.91, 2.92, 2.99, 2.100.

Medium: 2.17, 2.18, 2.24, 2.26, 2.34, 2.35, 2.36, 2.49, 2.50, 2.57, 2.58, 2.59, 2.60, 2.61, 2.63, 2.64, 2.65, 2.66, 2.67, 2.68, 2.69, 2.70, 2.71, 2.72, 2.73, 2.74, 2.75, 2.76, 2.79, 2.80, 2.81, 2.82, 2.83, 2.84, 2.85, 2.86, 2.88, 2.89, 2.90, 2.93, 2.94, 2.95, 2.96, 2.98, 2.101, 2.102, 2.103, 2.110, 2.112, 2.114, 2.115.

Difficult: 2.25, 2.77, 2.78, 2.87, 2.97, 2.104, 2.105, 2.106, 2.107, 2.108, 2.109, 2.111, 2.113.

First, convert 1 cm to picometers.

$$? \text{ He atoms} = (1 \times 10^{10} \cancel{\text{ pm}}) \times \frac{1 \text{ He atom}}{1 \times 10^2 \cancel{\text{ pm}}} = 1 \times 10^8 \text{ He atoms}$$

Note that you are given information to set up the unit factor relating meters and miles.

$$r_{\text{atom}} = 10^4 r_{\text{nucleus}} = 10^4 \times 2.0 \cancel{\text{ cm}} \times \frac{1 \cancel{\text{ m}}}{100 \cancel{\text{ cm}}} \times \frac{1 \text{ mi}}{1609 \cancel{\text{ m}}} = 0.12 \text{ mi}$$

For iron, the atomic number Z is 26. Therefore the mass number A is:

$$A = 26 + 28 = 54$$

Strategy: The 239 in Pu-239 is the mass number. The **mass number** (A) is the total number of

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**Thirumalaisamy P.
Velavancorresponding**



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Table of Contents Chang Solution Manual For

1. Understanding the eBook Chang Solution Manual For
 - The Rise of Digital Reading Chang Solution Manual For
 - Advantages of eBooks Over Traditional Books
2. Identifying Chang Solution Manual For
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chang Solution Manual For
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chang Solution Manual For
 - Personalized Recommendations
 - Chang Solution Manual For User Reviews and Ratings
 - Chang Solution Manual For and Bestseller Lists
5. Accessing Chang Solution Manual For Free and Paid eBooks
 - Chang Solution Manual For Public Domain eBooks
 - Chang Solution Manual For eBook Subscription Services
 - Chang Solution Manual For Budget-Friendly Options

6. Navigating Chang Solution Manual For eBook Formats
 - ePub, PDF, MOBI, and More
 - Chang Solution Manual For Compatibility with Devices
 - Chang Solution Manual For Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chang Solution Manual For
 - Highlighting and Note-Taking Chang Solution Manual For
 - Interactive Elements Chang Solution Manual For
8. Staying Engaged with Chang Solution Manual For
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chang Solution Manual For
9. Balancing eBooks and Physical Books Chang Solution Manual For
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chang Solution Manual For
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chang Solution Manual For
 - Setting Reading Goals Chang Solution Manual For
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chang Solution Manual For
 - Fact-Checking eBook Content of Chang Solution Manual For
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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