CHAPTER 10 The Mole

Section 1 Measuring Matter

page 320-324

Practice Problems

pages 323-324

 Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

-1.51×1026 atoms of Zn

 Calculate the number of molecules in 11.5 mol of water (H₂O).

= 6.92×10³⁴ molecules H₂O

 Silver nitrate (AgNO₁) is used to make several different silver halides used in photographic films.
How many formula units of AgNO₃ are there in 3.25 mol AgNO₁?

3.25 mol AgNO,
$$\times \frac{6.02 \times 10^{10} \text{ formula units AgNO,}}{1 \text{ mol AgNO,}}$$

= 1.96 × 1014 formula units AgNO,

 Challenge Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule. Ox.

5.00 mol
$$O_2 \times \frac{6.02 \times 10^{21} \text{ molecules } O_2}{1 \text{ mol} O_2}$$

- 5. How many moles contain each of the following?
 - 5.75×10²⁴ atoms Al

$$5.75 \times 10^{14}$$
 atoms AI $\times \frac{\text{I mol AI}}{6.02 \times 10^{19}}$ atoms AI = 9.55 mol AI

b. 2.50×10²⁰ atoms Fe

Chemistry: Matter and Change

- Challenge Identify the representative particle for each formula and convert the given number of representative particles to moles.
 - a. 3.75×10²⁴ CO₂

The representative particle is a molecule.

3.75×10²⁴ molecules CO₂

 $\times \frac{1 \text{ mol CO}_2}{6.02 \times 10^{23} \text{ molecules CO}_2} = 6.23 \text{ mol CO}_2$

b. 3.58×10²³ ZnCh

The representative particle is a formula unit.

3.58×1021 formula units ZnCl,

× 1 mol ZnCl, 6.02×10¹³ formula units ZnCl,

0.595 mol ZnCl₂

Section 1 Review

page 324

Explain why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

 State the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number (6.02×10²³) of representative particles.

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Karen Timberlake

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