

Name: \_\_\_\_\_

## Balancing Chemical Equations

Directions: Balance each of the following chemical equations.



- 1)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_  $\text{K}_3\text{PO}_4$  + \_\_\_  $\text{H}_2\text{O}$
- 2)    \_\_\_  $\text{K}$  + \_\_\_  $\text{B}_2\text{O}_3$   $\rightarrow$  \_\_\_  $\text{K}_2\text{O}$  + \_\_\_  $\text{B}$
- 3)    \_\_\_  $\text{HCl}$  + \_\_\_  $\text{NaOH}$   $\rightarrow$  \_\_\_  $\text{NaCl}$  + \_\_\_  $\text{H}_2\text{O}$
- 4)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{NaNO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 5)    \_\_\_  $\text{C}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{CS}_2$
- 6)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 7)    \_\_\_  $\text{N}_2$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{N}_2\text{O}_5$
- 8)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{Mg}(\text{OH})_2$   $\rightarrow$  \_\_\_  $\text{Mg}_3(\text{PO}_4)_2$  + \_\_\_  $\text{H}_2\text{O}$
- 9)    \_\_\_  $\text{NaOH}$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{CO}_3$  + \_\_\_  $\text{H}_2\text{O}$
- 10)   \_\_\_  $\text{KOH}$  + \_\_\_  $\text{HBr}$   $\rightarrow$  \_\_\_  $\text{KBr}$  + \_\_\_  $\text{H}_2\text{O}$
- 11)   \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 12)   \_\_\_  $\text{Al}(\text{OH})_3$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Al}_2(\text{CO}_3)_3$  + \_\_\_  $\text{H}_2\text{O}$
- 13)   \_\_\_  $\text{Al}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{Al}_2\text{S}_3$
- 14)   \_\_\_  $\text{Cs}$  + \_\_\_  $\text{N}_2$   $\rightarrow$  \_\_\_  $\text{Cs}_3\text{N}$
- 15)   \_\_\_  $\text{Mg}$  + \_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_  $\text{MgCl}_2$
- 16)   \_\_\_  $\text{Rb}$  + \_\_\_  $\text{RbNO}_3$   $\rightarrow$  \_\_\_  $\text{Rb}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 17)   \_\_\_  $\text{C}_6\text{H}_6$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$
- 18)   \_\_\_  $\text{N}_2$  + \_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_  $\text{NH}_3$

# Chapter 10 Challenge Problems Balancing Chemical Equations Answers

**Peter Scott, Rachael Enriquez Porter**



## **Chapter 10 Challenge Problems Balancing Chemical Equations Answers:**

## **Chapter 10 Challenge Problems Balancing Chemical Equations Answers** Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Chapter 10 Challenge Problems Balancing Chemical Equations Answers**," written by a highly acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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