

## Chapter 10 Practice Problems & Test Review

### Chemical Quantities **Answer Sheet**

Ⓢ Show all work (no work, no credit) **Answers in RED**

1. How many moles is 750 billion atoms H in water? (1 billion =  $1 \times 10^9$ )

$$\text{H} = 1.25 \times 10^{-12} \text{ H moles} \quad \text{H}_2\text{O} = 6.25 \times 10^{-13} \text{ moles H}_2\text{O}$$

2. How many carbon atoms are in a 9 mole gasoline? (gasoline =  $\text{C}_8\text{H}_{18}$ )

$$4.32 \times 10^{25} \text{ C atoms}$$

3. What is the molar mass of Iron (III) Oxalate?

$$\text{Fe}_2(\text{C}_2\text{O}_4)_3 = 375.6 \text{ g/mol}$$

4. How many moles are 500 grams Iron (III) Oxalate?

$$1.33 \text{ mol Fe}_2(\text{C}_2\text{O}_4)_3$$

5. How many grams is 1.78 moles Iron (II) Hydroxide?

$$\text{Fe}(\text{OH})_2 = 89.8 \text{ g/mol} \quad 159.8 \text{ g Fe}(\text{OH})_2$$

6. Calculate the volume of 3.44 moles methane,  $\text{CH}_4$ , gas at STP:

$$77.06 \text{ liters CH}_4$$

7. What is the density of Nitrogen Dioxide at STP?

$$\text{NO}_2 \quad 46 \text{ g/mol} \quad 2.05 \text{ g/l NO}_2$$

8. What is the volume of  $3.24 \times 10^{22}$  molecules of Chlorine gas at STP?

$$1.21 \text{ liters Cl}_2$$

9. How many H atoms and O atoms are in a 500 ml water bottle?

$$\text{H atoms} = 3.33 \times 10^{25}$$

$$\text{O atoms} = 1.67 \times 10^{25}$$

10. A dime is an alloy (a mixture of two solids) of 75% Copper (Cu) and 25% Nickel (Ni) and has a mass of 2.27 grams; how many Cu, Ni, and total atoms does a dime contain?

$$\text{Cu} = 1.61 \times 10^{22}$$

$$\text{Ni} = 5.80 \times 10^{21}$$

$$\text{total atoms} = 2.19 \times 10^{22}$$

Given a container with 50-ml of Ethanol ( $\text{C}_2\text{H}_6\text{O}$ ); the density of Ethanol is 0.789g/ml

11. Ethanol's molar mass 46 g/mol

12. Moles of ethanol 1.166 moles Ethanol

13. Carbon atoms  $1.4 \times 10^{24}$  C atoms

# Chapter 10 Chemical Quantities Practice Problems Answers

**Heather Hattori, Richard H. Langley**



## Chapter 10 Chemical Quantities Practice Problems Answers:

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*Annals of Clinical Biochemistry*, 1982

**The Practice of Chemistry** Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink [http://www.bfwpub.com/pdfs/winkPOCPowerPoint\\_Final.ppt](http://www.bfwpub.com/pdfs/winkPOCPowerPoint_Final.ppt) 832KB

**Problems for Introductory University Chemistry** James Newton Butler, Basil A. Dunell, Lionel G. Harrison, 1967

**Schaum's Outline of Theory and Problems of Chemistry Foundations** David E. Goldberg, 1991 A solved problem Outline for students with no high school chemistry background or a poor course in high school chemistry All topics are presented at an elementary level to commensurate with the introductory course simpler math is assumed developed a less intense approach to the basic concepts of chemistry a simpler manner of presentation There are hundreds of solved supplementary problems

*Chemical Engineering Catalog*, 1925

*Study Guide with Selected Solutions* Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

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[Introduction to Chemical Principles](#) Howard Stephen Stoker, 1990 Introduction to Chemical Principles is a text for students who have had little to no previous instruction in chemistry or who had such instruction long enough ago that a thorough review is needed preface

[Journal of the Society of Chemical Industry](#) Society of Chemical Industry (Great Britain), 1925 Includes list of members 1882 1902 proceedings of the annual meetings and various supplements

**The Journal of the Chemical, Metallurgical & Mining Society of South Africa**, 1926

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