

CHEMICAL QUANTITIES

Practice Problems

On a piece of paper, solve the following problems.

YOU MUST SHOW YOUR WORK TO GET CREDIT FOR YOUR ANSWERS!!!

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER

- What is the molar mass of sucrose (C₁₂H₂₂O₁₁)?
- 2. What is the molar mass of each of the following compounds?
 - a, phosphorus pentachloride (PCI-)
 - b. uranium hexafluoride (UF₄)
- 3. Calculate the molar mass of each of the following ionic compounds:
 - a. KMinO₄
 - b. Cas(PO₄):
- 4. How many moles is 3.52 × 1024 molecules of water?
- 5. How many atoms of zinc are in 0.60 mol of zinc?
- 6. What is the mass of 1.00 mol of exygen (O2)?

SECTION 10.2 MOLE-MASS AND MOLE-VOLUME RELATIONSHIPS

- 1. What is the molar mass of each of the following compounds?
 - a. CaHagOa

b. NaHCOs.

e. Calling

d. KNH4SO₄

- 2. Calculate the mass in grams of each of the following:
 - a. 8.0 mol lead oxide (PbO).

- d. 1.50 × 10⁻² mol molecular oxygen (O₂)
- b. 0.75 mol hydrogen sulfide (H₂S)

- e. 2.30 mol ethylene glycol (C₂H₂O₂)
- e. 0.00100 mol silicon tetrahydride (SiH₄).
- How many grams are in 1.73 mol of dinitrogen pentoxide (N₂O₂)?
- How many grams are in 0.658 mol of calcium phosphate [Ca₂(PO₄)₂]?
- 5. Calculate the number of moles in each of the following:
 - a. 0.50 g sodium bromide (NaBr).

- d, 0.00100 g monochloromethane (CH₂CI)
- b. 13.5 g magnesium nitrate [Mg(NO₁)₂]
- e. 1.50×10^{-3} g propylene glycol [C₃H₄(OH)₂]
- e. 1.02 g magnesium chloride (MgCl₂)
- 6. A chemist plans to use 435.0 grams of ammonium nitrate (NH₄NO₅) in a reaction. How many moles of the compound is this?
- 7. A solution is to be prepared in a laboratory. The solution requires 0.0465 mol of quinine (C₂₀H₂₄N₂O₂). What mass, in grams, should the laboratory technician obtain in order to make the solution?
- 8. What is the volume at STP of 2.66 mol of methane (CHa) gas?
- 9. How many moles is 135 L of ammonia (NH₃) gas at STP?

Chapter 10 Study Guide The Mole Answers

Iris Stovall

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