

Alcohol reacts with metallic Na to form sodium alkoxide with the liberation of H_2 .
 $ROH + Na \rightarrow RNa + 1/2 H_2$

Question 2.

Ethyl alcohol gives iodoform test. Methyl alcohol does not give iodoform test.

1. Do you agree with this?
2. If yes or no, substantiate your view.
3. How can you distinguish between 1 -butanol and 2-butanol?

Answer:

1. Yes.

2. Compounds containing CH_3CO - group and $CH_3CH(OH)$ - group on reaction with iodine and alkali give yellow colour of iodoform. Ethanol contains $CH_3CH(OH)$ - group.

3. 2-Butanol gives iodoform test as it contains $CH_3CH(OH)$ -group, whereas 1-Butanol does not answer iodoform test.

Question 3.

When an organic compound is treated with neutral ferric chloride a violet colour is obtained. What will be the compound? Explain.

Answer:

Phenol. Phenol forms a violet-coloured water soluble complex with ferric chloride.

Question 4.

- O - is the functional group of ether. Classify the following into two groups with an appropriate heading.

CH_3-O-CH_3 , $CH_3-O-C_2H_5$, $C_2H_5-O-C_2H_5$, $C_6H_5-O-CH_3$.

Answer:

Symmetric ether	Unsymmetric ether
$CH_3 - O - CH_3$ $C_2H_5 - O - C_2H_5$	$CH_3 - O - C_2H_5$  $O - CH_3$

Question 5.

Explain how does the -OH group attached to a carbon of benzene ring activates it towards electrophilic substitution reaction?

Answer:

In phenol, the -OH group is directly attached to a carbon of benzene ring. The lone pair of oxygen participates in resonance with the benzene ring. As a result, electron

Chapter 11 Chemistry Answers

H Kauffman



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Table of Contents Chapter 11 Chemistry Answers

1. Understanding the eBook Chapter 11 Chemistry Answers
 - The Rise of Digital Reading Chapter 11 Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 11 Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Chemistry Answers
 - Personalized Recommendations

- Chapter 11 Chemistry Answers User Reviews and Ratings
- Chapter 11 Chemistry Answers and Bestseller Lists
- 5. Accessing Chapter 11 Chemistry Answers Free and Paid eBooks
 - Chapter 11 Chemistry Answers Public Domain eBooks
 - Chapter 11 Chemistry Answers eBook Subscription Services
 - Chapter 11 Chemistry Answers Budget-Friendly Options
- 6. Navigating Chapter 11 Chemistry Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 11 Chemistry Answers Compatibility with Devices
 - Chapter 11 Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 11 Chemistry Answers
 - Highlighting and Note-Taking Chapter 11 Chemistry Answers
 - Interactive Elements Chapter 11 Chemistry Answers
- 8. Staying Engaged with Chapter 11 Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 11 Chemistry Answers
- 9. Balancing eBooks and Physical Books Chapter 11 Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 11 Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 11 Chemistry Answers
 - Setting Reading Goals Chapter 11 Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 11 Chemistry Answers
 - Fact-Checking eBook Content of Chapter 11 Chemistry Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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