Name _____ Date _____ Class ____

CHAPTER 11, Chemical Reactions (continued)

 Use the symbols in Table 11.1 on page 323 to write a skeleton equation for the following chemical reaction. Hydrochloric acid reacts with zinc to produce aqueous zinc(II) chloride and hydrogen gas.

 $HCI(aq) + Zn(s) \longrightarrow ZnCI_{r}(aq) + H_{r}(g)$

▶ Balancing Chemical Equations (pages 324–328)

10. What is the law of conservation of mass?

In any physical or chemical change, mass is neither created nor destroyed.

11. Complete the flowchart for balancing equations.

Determine the correct formulas and physical states for

the reactants and products

Write a <u>skeleton equation</u> with the formulas for the reactants on the left and the formulas for the products on the right of a yields sign (———).

Count the number of ______ of each element in the reactants and in the products.

Check each atom or polyatomic ion to be sure the equation is ______, and make sure that all coefficients are in the ______ possible ratio.

Chapter 11 Describing Chemical Reactions Answers

Arthur Horseman Hiorns

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