

Molecular Composition of Gases

SECTION: PART B (20)

QUESTION: ANSWER Answer the following questions in the space provided.

1. _____ The average speed of a gas molecule is most directly related to the _____.
 - (a) volume of the container
 - (b) pressure within the gas
 - (c) temperature of the gas
 - (d) number of molecules in the gas
2. _____ Which value below best represents the average molecular speed of a diatomic gas given at the same temperature?
 - (a) The value associated with molecular mass is directly proportional.
 - (b) The value associated with molecular mass is directly proportional.
 - (c) The value associated with square root of molecular mass is directly proportional.
 - (d) The value associated with the square of molecular mass is directly proportional.
3. _____ Which combination of conditions corresponds to the following equation:

$$\text{volume} \propto \text{pressure} \propto \text{temperature}$$

 - a. _____ $\text{PV} = nRT$ is correct. The volume of H_2 is constant and the pressure of H_2 is constant.
 - b. _____ $\text{PV} = nRT$ is correct. The volume of H_2 is constant and the pressure of H_2 is constant.
4. _____ Which set of conditions is not a possible set of values for a gas?
 - a. _____ Mass, volume, molecular pressure, the weight of the gas is constant.
 - b. _____ Mass, volume, molecular pressure, the weight of the gas is constant.
 - c. _____ Molecular pressure, the weight of the gas is constant.

PROBLEM: ANSWER The gas that occupies less than three has the least mass. Nitrogen will "grow" under the same conditions.

5. _____ Calculate the volume of 1 mole of gas at 1 atm and 273 K.

Chapter 11 Review Molecular Composition Of Gases

Mixed Answers

C Cleary



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 **Environmental Science** Scott Brennan, 2004-03 This comprehensive study guide features chapter outlines key terms practice exercises and answers and solutions to selected end of chapter questions from the text Additional suggestions for approaching environmental issues encourage students to think critically **Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03

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