

Physics

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(Chapter 11) (Thermodynamics) (Exercises)

(Class 11)

Question 11.1:

A geyser heats water flowing at the rate of 3.0 litres per minute from 27 °C to 77 °C. If the geyser operates on a gas burner, what is the rate of consumption of the fuel if its heat of combustion is 4.0×10^4 J/g?

Answer 11.1:

Water is flowing at a rate of 3.0 litre/min.

The geyser heats the water, raising the temperature from 27°C to 77°C.

Initial temperature, $T_1 = 27^\circ\text{C}$ and Final temperature, $T_2 = 77^\circ\text{C}$

Rise in temperature, $\Delta T = T_2 - T_1 = 77 - 27 = 50^\circ\text{C}$

Heat of combustion = 4×10^4 J/g

Specific heat of water, $c = 4.2 \text{ J g}^{-1} \text{ }^\circ\text{C}^{-1}$

Mass of flowing water, $m = 3.0 \text{ litre/min} = 3000 \text{ g/min}$

Total heat used, $\Delta Q = mc \Delta T = 3000 \times 4.2 \times 50 = 6.3 \times 10^5 \text{ J/min}$

\therefore Rate of consumption = $\frac{6.3 \times 10^5}{4 \times 10^4} = 15.75 \text{ g/min}$

Question 11.2:

What amount of heat must be supplied to 2.0×10^{-2} kg of nitrogen (at room temperature) to raise its temperature by 45 °C at constant pressure? (Molecular mass of $N_2 = 28$; $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$.)

Answer 11.2:

Mass of nitrogen, $m = 2.0 \times 10^{-2} \text{ kg} = 20 \text{ g}$ and Rise in temperature, $\Delta T = 45^\circ\text{C}$

Molecular mass of N_2 , $M = 28$

Universal gas constant, $R = 8.3 \text{ J mol}^{-1} \text{ K}^{-1}$

Number of moles, $n = \frac{m}{M} = \frac{2.0 \times 10^{-2} \times 10^3}{28} = 0.714$

Molar specific heat at constant pressure for nitrogen $C_p = \frac{7}{2}R = \frac{7}{2} \times 8.3 = 29.05 \text{ J mol}^{-1} \text{ K}^{-1}$

The total amount of heat to be supplied is given by the relation: $\Delta Q = nC_p \Delta T$

$= 0.714 \times 29.05 \times 45 = 933.38 \text{ J}$

Therefore, the amount of heat to be supplied is 933.38 J.

Question 11.3:

Explain why

- (a) Two bodies at different temperatures T_1 and T_2 if brought in thermal contact do not necessarily settle to the mean temperature $(T_1 + T_2)/2$.
- (b) The coolant in a chemical or a nuclear plant (i.e., the liquid used to prevent the different parts of a plant from getting too hot) should have high specific heat.
- (c) Air pressure in a car tyre increases during driving.
- (d) The climate of a harbour town is more temperate than that of a town in a desert at the same latitude.

Answer 11.3:

- (a) When two bodies at different temperatures T_1 and T_2 are brought in thermal contact, heat flows from the body at the higher temperature to the body at the lower temperature till equilibrium is achieved, i.e., the temperatures of both the bodies become equal. The equilibrium temperature is equal to the mean temperature $(T_1 + T_2)/2$ only when the thermal capacities of both the bodies are equal.

Chapter 11 Solutions Thermodynamics An Engineering Approach

David Kirk



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Introduction to Chemical Engineering Thermodynamics Joseph Mauk Smith, Hendrick C. Van Ness, Michael M. Abbott, 1996 Introduction to Chemical Engineering Thermodynamics Fifth Edition presents a thorough exposition of the principles of thermodynamics and details their application to chemical processes Newly revised and completely up to date this best selling book also equips the reader with an adequate foundation for subsequent self instruction Learner friendly the fifth edition of Introduction to Chemical Engineering Thermodynamics includes over 115 worked examples as well as 8 helpful appendices This classic textbook is written not only for students but also for practicing engineers

Introduction to Engineering Fluid Mechanics Marcel Escudier, 2017 Fluid mechanics concerns the way fluids flow in response to imposed stresses This textbook includes numerous examples of practical applications of the theoretical ideas such as calculations of the thrust of a jet engine the power output of a gas turbine and forces created by liquid flow through a pipe bend or junction

Introductory Chemical Engineering Thermodynamics J. Richard Elliott, Carl T. Lira, 2012 In this book two leading experts and long time instructors thoroughly explain thermodynamics taking the molecular perspective that working engineers require This edition contains extensive new coverage of today's fast growing biochemical engineering applications notably biomass conversion to fuels and chemicals It also presents many new MATLAB examples and tools to complement its previous usage of Excel and other software

Applied Mechanics Reviews, 1962

NASA Tech Briefs, 1991

Advanced Mathematics in Physics and Engineering Arthur B. Bronwell, 1953

Inelasticity of Materials Arun Ramaswamy Srinivasa, Srinivasan M. Sivakumar, 2009 With the advent of a host of new materials ranging from shape memory alloys to biomaterials to multiphase alloys acquiring the capacity to model inelastic behavior and to choose the right model in a commercial analysis software has become a pressing need for practicing engineers Even with the traditional materials there is a continued emphasis on optimizing and extending their full range of capability in the applications This textbook builds upon the existing knowledge of elasticity and thermodynamics and allows the reader to gain confidence in extending one's skills in understanding and analyzing problems in inelasticity By reading this textbook and working through the assigned exercises the reader will gain a level of comfort and competence in developing and using inelasticity models Thus the book serves as a valuable book for practicing engineers and senior level undergraduate graduate level students in the mechanical civil aeronautical metallurgical and other disciplines The book is written in three parts Part 1 is primarily focused on lumped parameter models and simple structural elements such as trusses and beams This is suitable for an advanced undergraduate class with just a strength of materials background Part II is focused on small deformation multi dimensional inelasticity and is suitable for a beginning graduate class Sufficient material is included on how to numerically implement an inelastic model and solve either using a simple stress function type of approach or using commercial software Case studies are included as examples There is also an extensive discussion of thermodynamics in the context of small deformations Part

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Thermodynamics: an Introduction

Rubin Battino, Scott Emerson Wood, 1968 Introduction to Chemical Engineering and Computer Calculations Alan L. Myers, Warren D. Seider, 1976 *Fox and McDonald's Introduction to Fluid Mechanics* Robert W. Fox, Alan T.

McDonald, John W. Mitchell, 2020-06-30 Through ten editions Fox and McDonald's Introduction to Fluid Mechanics has helped students understand the physical concepts basic principles and analysis methods of fluid mechanics This market leading textbook provides a balanced systematic approach to mastering critical concepts with the proven Fox McDonald solution methodology In depth yet accessible chapters present governing equations clearly state assumptions and relate mathematical results to corresponding physical behavior Emphasis is placed on the use of control volumes to support a practical theoretically inclusive problem solving approach to the subject Each comprehensive chapter includes numerous easy to follow examples that illustrate good solution technique and explain challenging points A broad range of carefully selected topics describe how to apply the governing equations to various problems and explain physical concepts to enable students to model real world fluid flow situations Topics include flow measurement dimensional analysis and similitude flow in pipes ducts and open channels fluid machinery and more To enhance student learning the book incorporates numerous pedagogical features including chapter summaries and learning objectives end of chapter problems useful equations and design and open ended problems that encourage students to apply fluid mechanics principles to the design of devices and systems *Chemical Principles* Steven S. Zumdahl, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

Mathematical Reviews ,2005 **Chemical Engineering Progress** ,2000 *The Engineer* ,1885 **Journal of the**

Indian Chemical Society Indian Chemical Society,1997 **Directory of Graduate Research** American Chemical Society.
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