Solution Concentrations REVIEW

NAME:_____

Solution → Solute (KoolAid) + Solvent (Water)

Percent by Mass =
$$\frac{Mass \text{ of } Solute(g)}{Mass \text{ of } Solution(g)} \times 100$$

Percent by Volume = $\frac{Volume \text{ of } Solute(mL)}{Volume \text{ of } Solution(mL)} \times 100$

Mass | Volume = $\frac{Mass \text{ of } solute(g)}{Volume \text{ of } Solution(mL)}$

 $Molarity(M) = \frac{Moles of solute}{Volume of Solution i.i.}$

 Glucose is a sugar that is found abundantly in nature. What is the percent by mass of a solution made by dissolving 163 g of glucose in 755 g of water? Do you need to know the formula of glucose? Why or why not?

2. What is the mass percent sucrose in a solution obtained by mixing 225 g of sucrose with 135 g water?

Determine the volume percent of toluene in a solution made by mixing 40.0 mL toluene with 75.0 mL of benzene.

 What is the molarity of a solution prepared by dissolving 2.25g of glucose (C6H12O6) in 0.500 L. solution.

Concentrations Of Solutions Section Review Answers

Philip C. Keller, Raymond Chang, Jill L. Keller

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