

(Group – I)

Note: Attempt any FIFTEEN (15) questions from Part – I and attempt any TWO questions from Part – II.

PART – I

2. Write any FIFTEEN (15) short answers of the following questions: **30**

- (i) Define hygroscopic compounds.
- (ii) Write two disadvantages of hard water.
- (iii) Write down names of polluted water born disease.
- (iv) What is olvine? Write its chemical formula.
- (v) Define semi-conductor.
- (vi) What is meant by oxidation? Describe with example.
- (vii) Write the names of different layers of atmosphere around the earth.
- (viii) Write down two uses of hydrogen peroxide.
- (ix) Why sulphuric acid behaves as dehydrating agent?
- (x) How can you prepare pure sample of sulphur from mixture of sand and sulphur?
- (xi) Define halogens and write the names of halogens.
- (xii) How chlorine is prepared on industrial scale?
- (xiii) What is Nascent Oxygen?
- (xiv) Write the four uses of methane.
- (xv) Write the chemical equation for the preparation of ethyne in laboratory.
- (xvi) Define homologous series.
- (xvii) Describe the two reasons of food spoilage.
- (xviii) Write the names of four types of soap.
- (xix) Write the names of four raw material for the preparation of stamp pad ink.
- (xx) Write the four uses of hydrochloric acid.
- (xxi) Why ice floats on water?
- (xxii) Define isomerism and give an example.

PART – II

- 3.(a)** Write reactions of hydrogen with the following: **4**
 (i) CuO (ii) P (iii) Na (iv) Ca
- (b)** Describe the structure and uses of diamond. **5**
- 4.(a)** Write physical and chemical properties along with two uses of ozone. **4**
- (b)** Give the reaction of sulphur with the following: **5**
 (i) H₂ (ii) Pb (iii) H₂SO₄ (iv) Cℓ₂ v) Carbon
- 5.(a)** Write down the balanced chemical equations for the reaction of bleaching powder with following: **4**
 (i) NH₃ (ii) HCℓ (iii) H₂SO₄ (iv) H₂O₂
- (b)** Write down five characteristics of organic compounds. **5**

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