

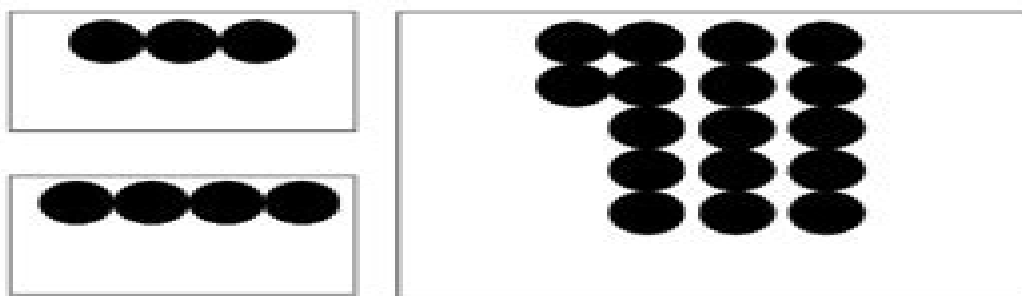
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Balancing Act

Why?

In a chemical reaction, matter cannot be created or destroyed. The mass of the reactants in a chemical reaction must be equal to the mass of the products. It is not possible to start a chemical reaction with 2 atoms of hydrogen and produce 2 atoms of gold. Similarly, it is not possible to start a chemical reaction with 2 atoms of hydrogen and produce only 1 atom of hydrogen. When a chemical equation is written, the same number of each type of atom must appear on both the product and reactant side.

Model 1



1) According to Model 1:

a) How many atoms of hydrogen and oxygen are in 1 molecule of water, H_2O ?

H: 2 O: 1

b) How many atoms of hydrogen and oxygen are in 1 molecule of H_2O_2 , hydrogen peroxide?

H: 2 O: 2

c) How many atoms of aluminum, sulfur, and oxygen are in 1 molecule of aluminum sulfate, $\text{Al}_2(\text{SO}_4)_3$?

Al: 2 S: 3 O: 12

d) How many atoms of hydrogen and oxygen are in 2 molecules of water?

H: 4 O: 2

e) How many atoms of hydrogen and oxygen are in 8 molecules of hydrogen peroxide?

H: 16 O: 16

f) How many atoms of aluminum, sulfur, and oxygen are in 3 molecules of aluminum sulfate?

Al: 6 S: 9 O: 36

2) Write how would you indicate 2 water molecules using chemical symbols: $2\text{H}_2\text{O}$.

3) Using complete sentences, explain the difference between these two chemical formulas: $4\text{H}_2\text{O}$ and H_4O_4 .

Balancing Act Practice Answers

Eleanor Frances Jourdain

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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A Great Balancing Act Anne Chapman, 1997 Creating equitable education for boys and girls.

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