

Identifying Reaction Types & Balancing Equations

<u>Step 1:</u> Name the type of chemical reaction	<u>Step 2:</u> Balance the following chemical equations:
1. Ex: Synthesis	$\text{___ N}_2 + \text{___ H}_2 \rightarrow \text{___ NH}_3$
2.	$\text{___ KClO}_3 \rightarrow \text{___ KCl} + \text{___ O}_2$
3.	$\text{___ NaCl} + \text{___ F}_2 \rightarrow \text{___ NaF} + \text{___ Cl}_2$
4.	$\text{___ H}_2 + \text{___ O}_2 \rightarrow \text{___ H}_2\text{O}$
5.	$\text{___ AgNO}_3 + \text{___ MgCl}_2 \rightarrow \text{___ AgCl} + \text{___ Mg(NO}_3)_2$
6.	$\text{___ AlBr}_3 + \text{___ K}_2\text{SO}_4 \rightarrow \text{___ KBr} + \text{___ Al}_2(\text{SO}_4)_3$
7.	$\text{___ CH}_4 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
8.	$\text{___ C}_3\text{H}_8 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
9.	$\text{___ FeCl}_3 + \text{___ NaOH} \rightarrow \text{___ Fe(OH)}_3 + \text{___ NaCl}$
10.	$\text{___ P} + \text{___ O}_2 \rightarrow \text{___ P}_2\text{O}_5$
11.	$\text{___ Na} + \text{___ H}_2\text{O} \rightarrow \text{___ NaOH} + \text{___ H}_2$
12.	$\text{___ Ag}_2\text{O} \rightarrow \text{___ Ag} + \text{___ O}_2$
13.	$\text{___ CO}_2 + \text{___ H}_2\text{O} \rightarrow \text{___ C}_6\text{H}_{12}\text{O}_6 + \text{___ O}_2$
14.	$\text{___ K} + \text{___ MgBr}_2 \rightarrow \text{___ KBr} + \text{___ Mg}$
15.	$\text{___ HNO}_3(\text{aq}) + \text{___ Ba(OH)}_2(\text{aq}) \rightarrow \text{___ Ba(NO}_3)_2(\text{aq}) + \text{___ H}_2\text{O (l)}$
16.	$\text{___ C}_5\text{H}_{12}(\text{g}) + \text{___ O}_2(\text{g}) \rightarrow \text{___ CO}_2(\text{g}) + \text{___ H}_2\text{O(g)}$
17.	$\text{___ Al(s)} + \text{___ Fe}_2\text{O}_3(\text{aq}) \rightarrow \text{___ Al}_2\text{O}_3(\text{aq}) + \text{___ Fe(s)}$
18.	$\text{___ Al(s)} + \text{___ O}_2(\text{g}) \rightarrow \text{___ Al}_2\text{O}_3(\text{s})$

Balancing Equations And Identifying Reaction Types

Elizabeth Kean, Catherine Middlecamp



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