

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Derive the molecular formula:

1. The empirical formula of a compound is PCl_2 . Its molecular mass is 263 g/mol. What is its molecular formula?

Given: Empirical formula =

Mass

2. The empirical formula of a compound is CH_2 . Its molecular mass is 78 g/mol. What is its molecular formula?

Mass of 1 mol = 12 g/mol, Mass of 2 mol = 24 g/mol, Mass of 3 mol = 36 g/mol

Mass

3. A compound is found to contain 40.0% carbon, 5.15% hydrogen and 54.8% oxygen. Its molecular mass is 176 g/mol. What is its molecular formula?

Empirical formula = $\text{C}_4\text{H}_5\text{O}_2$

Empirical mass = 89 g/mol

Molecular mass = 176 g/mol

Mass

4. A compound is found to contain 52.0% carbon, 4.47% hydrogen and 43.5% oxygen. Its molecular mass is 106 g/mol. What is its molecular formula?

Empirical formula = $\text{C}_4\text{H}_4\text{O}_2$

Empirical mass = 88 g/mol

Molecular mass = 106 g/mol

Mass

5. A compound is found to contain 64.0% carbon, 5.62% hydrogen and 30.38% oxygen. Its molecular mass is 152 g/mol. What is its molecular formula?

Empirical formula = $\text{C}_4\text{H}_5\text{O}_2$

Empirical mass = 89 g/mol

Molecular mass = 152 g/mol

Mass

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Howard Stephen Stoker



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