

SC3-4 Atomic structure and periodic table exam questions

Q1.

The elements beryllium, magnesium, calcium, strontium and barium are in group 2 of the periodic table. There are three common isotopes of magnesium.

(i) These isotopes are shown in Figure 6.

Complete Figure 6 to show the number of protons and neutrons in an atom of each of the other two isotopes.

(2)

isotope	mass number	number of protons in an atom	number of neutrons in an atom
magnesium-24	24	12	12
magnesium-25	25		
magnesium-26	26		

Figure 6

(ii) A sample of magnesium contains 78.60% magnesium-24, 10.11% magnesium-25 and 11.29% magnesium-26.

Use this information to calculate the relative atomic mass of magnesium in this sample. Give your answer to 3 significant figures.

(4)

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relative atomic mass =

Q2.

In 1871, the Russian chemist Dmitri Mendeleev produced the first version of the periodic table.

Describe how Mendeleev arranged the elements in this first version of the periodic table and why the elements are arranged differently in the modern periodic table.

(3)

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(Total for question = 3 marks)

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