

MATHEMATICS STANDARD – Code No.041
SAMPLE QUESTION PAPER
CLASS – X (2025-26)

Maximum Marks: 80

Time: 3 hours

General Instructions:

Read the following instructions carefully and follow them:

1. This question paper contains 38 questions. All Questions are compulsory.
2. This Question Paper is divided into 5 Sections A, B, C, D and E.
3. In Section A, Question numbers 1-18 are multiple choice questions (MCQs) and questions no. 19 and 20 are Assertion- Reason based questions of 1 mark each.
4. In Section B, Question numbers 21-25 are very short answer (VSA) type questions, carrying 02 marks each.
5. In Section C, Question numbers 26-31 are short answer (SA) type questions, carrying 03 marks each.
6. In Section D, Question numbers 32-35 are long answer (LA) type questions, carrying 05 marks each.
7. In Section E, Question numbers 36-38 are case study-based questions carrying 4 marks each with sub parts of the values of 1, 1 and 2 marks each respectively.
8. There is no overall choice. However, an internal choice in 2 questions of Section B, 2 questions of Section C and 2 questions of Section D has been provided. An internal choice has been provided in all the 2 marks questions of Section E.

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web chemistry chemistry questions and answers lab 16 lewis dot structures molecular models vsepr theory valence bond theory and polarity name date formula number of lewis dot electrons formula molecular electron geometry drawing of structure hybridization geometry hình trigonal 8 tetrahedral sp³ pyramidal nh₃ h fb f

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web problem 4 2 7 4 2 7 the arrangement of atoms in several biologically important molecules is given here complete the lewis structures of these molecules by adding multiple bonds and lone pairs do not add any more atoms a the amino acid serine b urea c pyruvic acid d uracil e carbonic acid

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web introduction chemical structures there are a number of different ways to represent molecules as illustrated for NH_3 in figure 1 a lewis structure figure 1a is meant to represent chemical bonding and the distribution of electrons

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web nov 17 2023 introduction lewis structure is basically a graphic representation of the electron distribution around an atom the major reason why learning lewis dot structure is important is that it helps in predicting the number and type of bonds which can be formed around an atom it also helps in predicting the geometry of the molecule

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web nov 14 2023 show all resonance structures possible use formal charges to pick the best form if necessary the following molecule has two lewis dot formulas write one that uses 8 electrons about the central atom and one that expands the natat
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web 2019 0501140012 american native medicine 1 2019 0612171451 orpic company 1 biome viewer student ws copy of lab 1 effect environment se the periodic properties of the main group to understand theory behind lewis structure and valence shell electron pair repulsion theory to experience molecule s geometrical shapes by the model that has

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web mar 27 2023 a the lewis structure for water is represented by the given option a in this structure the oxygen atom is in the center and is bonded to two hydrogen atoms the oxygen atom has two lone pairs of electrons this structure follows the octet rule where each atom has a full valence shell with eight electrons

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web introduction the physical and chemical behavior of compounds depends on the types of bonds that exist within the molecule and on the overall shape of the molecules the branch of chemistry that investigates how molecules are put together is

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web lewis structures represent the valence electrons present in the molecule as lone pairs unshared electrons drawn as dots or bonds shared electrons drawn as lines procedure for drawing lewis structures add up the total number of valence electrons that each atom contributes to the molecule polyatomic ion

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web the pcl 3 lewis structure has the typical case of phosphorus p in the center with 3 bonds to 3 other atoms phosphorus is from the same column as nitrogen in the periodic table meaning that p and n generally have the same bonding structure note the lone pair dots without bonds on top of p just like for n in the previous example for nh 3 chlorine cl

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web step 1 figure out how many electrons the molecule must have based on the number of valence electrons in each atom when drawing the structure of an ion be sure to add subtract electrons to account for the charge step 2 connect the atoms to each other with single bonds to form a skeleton structure

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2 total number of valence electrons 24 cad engineered 3d sketch model show dipole arrows lewis structure is there a polar
bond in this molecule yes or no vsepr shape name trigonal planar bond angles 120

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web jun 8 2022 lewis structure in chemistry lewis structures also called electron dot structures or electron dot diagrams are
diagrams that show the bonding between atoms of a molecule and the lone pairs of

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web this exercise required us to take the formulas in the data table and calculate the number of valence electrons using the
periodic table draw its lewis structure on paper build the molecular model and record a photo list the number of bonds and
lone pairs surrounding the central atom and lastly identify the structure geometry