

CH101 Experiment 1: Preparing Solutions

Outcomes

Upon completion of this lab, the student should be able to

- proficiently calculate molarities for solutions.
- prepare a solution of known concentration.
- prepare a dilute solution from a more concentrated one.
- perform serial dilutions.
- use volumetric and Mohr pipets and a volumetric flask.
- perform calculations using errors

Background

Solutions are an important part of chemistry. In this lab you will practice preparing solutions of different concentrations. The amount of solute that is dissolved in a given quantity of solvent is expressed as the concentration of the solution. A dilute solution contains only a small amount of solute in a given amount of solution. The units chemists use most often to describe concentration of solutions are molarity units. The molarity, *M*, (sometimes denoted as Concentration "C") of a solution is the number of moles of solute in one litre of solution. To determine the molarity of a solution, the following equation can be used:

$$\text{Molarity (M)} = \frac{\text{moles of solute}}{\text{Litres of solution}}$$

Example 1: How would 500.0 mL of a 0.6000 M NaCl solution be prepared?

Since NaCl is a solid, its mass must be measured on a balance. The first step is to calculate the mass of solute needed. The molar mass for NaCl is 58.44 g/mol.

$$500.0 \text{ mL solution} \times \frac{1 \text{ L Solution}}{1000\text{-mL}} \times \frac{0.6000 \text{ mol NaCl}}{1 \text{ L Solution}} \times \frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} = 17.53 \text{ g NaCl}$$

Answer: Using a balance, measure out 17.53 g of NaCl into a small beaker and transfer it to a 500 -mL volumetric flask. A funnel may be necessary to transfer the solid into the flask. Rinse the beaker and funnel with a small amount of distilled water into the flask. Then add distilled water to the flask until it is about $\frac{3}{4}$ full. Cap the flask and invert it several times to dissolve the solid. Next, add water until the liquid is just below the etched line on the neck of the flask. Bring the water to the line by adding the last few drops of water drop-by-drop using either a wash bottle or dropper. Again, cap and invert the flask several times to ensure proper mixing. The final volume of the solution is 500.0 mL. See Figure 1 on the next page.

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Thomas Harris (of Baltimore.)



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