

4. a. 
$$\frac{10 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 35.33 \text{ g Cu}$$

$$\frac{20 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol Cu}} = \boxed{9.45 \text{ g Cu}}$$

b. 
$$\frac{30 \text{ g Al} \mid 1 \text{ mol Al} \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol AlCl}_3} = 148.25 \text{ g AlCl}_3$$

$$\frac{100 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = \boxed{66.11 \text{ g AlCl}_3}$$

c. 
$$\frac{5 \text{ g Al} \mid 1 \text{ mol Al} \mid 3 \text{ mol Cu} \mid 63.55 \text{ g Cu}}{26.98 \text{ g Al} \mid 2 \text{ mol Al} \mid 1 \text{ mol Cu}} = 17.67 \text{ g Cu}$$

~~100 g / 17.67 g~~  $10 \text{ g} / 17.67 \text{ g} \times 100 = \boxed{56.6\%}$

d. 
$$\frac{10 \text{ g CuCl}_2 \mid 1 \text{ mol CuCl}_2 \mid 2 \text{ mol AlCl}_3 \mid 133.33 \text{ g AlCl}_3}{134.45 \text{ g CuCl}_2 \mid 3 \text{ mol CuCl}_2 \mid 1 \text{ mol AlCl}_3} = 6.61 \text{ g AlCl}_3$$

~~100 g / 6.61 g~~  $5 \text{ g} / 6.61 \text{ g} \times 100 = \boxed{75.6\%}$

- 5.
- a. TRUE
  - b. FALSE
  - c.  $\text{AgNO}_3 \Rightarrow 169.91 \text{ g/mol}$
  - d.  $6.02 \times 10^{23}$  particles = 1 mole of particles

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**James E. Brady**



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1. Understanding the eBook Chemistry Study Guide Answers Stoichiometry
  - The Rise of Digital Reading Chemistry Study Guide Answers Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Study Guide Answers Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Study Guide Answers Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Study Guide Answers Stoichiometry
  - Personalized Recommendations
  - Chemistry Study Guide Answers Stoichiometry User Reviews and Ratings
  - Chemistry Study Guide Answers Stoichiometry and Bestseller Lists
5. Accessing Chemistry Study Guide Answers Stoichiometry Free and Paid eBooks
  - Chemistry Study Guide Answers Stoichiometry Public Domain eBooks
  - Chemistry Study Guide Answers Stoichiometry eBook Subscription Services
  - Chemistry Study Guide Answers Stoichiometry Budget-Friendly Options

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  - ePub, PDF, MOBI, and More
  - Chemistry Study Guide Answers Stoichiometry Compatibility with Devices
  - Chemistry Study Guide Answers Stoichiometry Enhanced eBook Features
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  - Adjustable Fonts and Text Sizes of Chemistry Study Guide Answers Stoichiometry
  - Highlighting and Note-Taking Chemistry Study Guide Answers Stoichiometry
  - Interactive Elements Chemistry Study Guide Answers Stoichiometry
8. Staying Engaged with Chemistry Study Guide Answers Stoichiometry
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Study Guide Answers Stoichiometry
9. Balancing eBooks and Physical Books Chemistry Study Guide Answers Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Study Guide Answers Stoichiometry
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Study Guide Answers Stoichiometry
  - Setting Reading Goals Chemistry Study Guide Answers Stoichiometry
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Study Guide Answers Stoichiometry
  - Fact-Checking eBook Content of Chemistry Study Guide Answers Stoichiometry
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
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14. Embracing eBook Trends
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