

## Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

- Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before  $\underline{2 \text{ mol}}$   $\underline{0 \text{ mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-2 \text{ mol}}$   $\underline{+2 \text{ mol}}$  \_\_\_\_\_

After  $\underline{0 \text{ mol}}$   $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

- Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before  $\underline{9.94 \text{ mol}}$   $\underline{\text{xs mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-9.94 \text{ mol}}$  \_\_\_\_\_  $\underline{+9.94 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

- You eat 180.0 g of glucose (90 M&Ms). If glucose,  $\text{C}_6\text{H}_{12}\text{O}_6$ , reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before  $\underline{1.0 \text{ mol}}$   $\underline{\text{xs mol}}$   $\underline{0 \text{ mol}}$   $\underline{0 \text{ mol}}$

Change  $\underline{-1.0 \text{ mol}}$  \_\_\_\_\_  $\underline{-1.0 \text{ mol}}$

After  $\underline{0 \text{ mol}}$   $\underline{\text{xs mol}}$

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

# Chemistry Supplemental Stoichiometry Problems Answer Key

**Bowker Editorial Staff**



**Chemistry Supplemental Stoichiometry Problems Answer Key:**

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### **Table of Contents Chemistry Supplemental Stoichiometry Problems Answer Key**

1. Understanding the eBook Chemistry Supplemental Stoichiometry Problems Answer Key
  - The Rise of Digital Reading Chemistry Supplemental Stoichiometry Problems Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Supplemental Stoichiometry Problems Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chemistry Supplemental Stoichiometry Problems Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Supplemental Stoichiometry Problems Answer Key
  - Personalized Recommendations

- Chemistry Supplemental Stoichiometry Problems Answer Key User Reviews and Ratings
- Chemistry Supplemental Stoichiometry Problems Answer Key and Bestseller Lists
- 5. Accessing Chemistry Supplemental Stoichiometry Problems Answer Key Free and Paid eBooks
  - Chemistry Supplemental Stoichiometry Problems Answer Key Public Domain eBooks
  - Chemistry Supplemental Stoichiometry Problems Answer Key eBook Subscription Services
  - Chemistry Supplemental Stoichiometry Problems Answer Key Budget-Friendly Options
- 6. Navigating Chemistry Supplemental Stoichiometry Problems Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chemistry Supplemental Stoichiometry Problems Answer Key Compatibility with Devices
  - Chemistry Supplemental Stoichiometry Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chemistry Supplemental Stoichiometry Problems Answer Key
  - Highlighting and Note-Taking Chemistry Supplemental Stoichiometry Problems Answer Key
  - Interactive Elements Chemistry Supplemental Stoichiometry Problems Answer Key
- 8. Staying Engaged with Chemistry Supplemental Stoichiometry Problems Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chemistry Supplemental Stoichiometry Problems Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Supplemental Stoichiometry Problems Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chemistry Supplemental Stoichiometry Problems Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Supplemental Stoichiometry Problems Answer Key
  - Setting Reading Goals Chemistry Supplemental Stoichiometry Problems Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Supplemental Stoichiometry Problems Answer Key
  - Fact-Checking eBook Content of Chemistry Supplemental Stoichiometry Problems Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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