

## Unit 5 – Counting Particles - Objectives

<p>Review Concepts</p> <ol style="list-style-type: none"> <li>Types of substances</li> <li>Chemical formulas of substances (U 4)</li> </ol>	
1. State evidence for Avogadro's Hypothesis. Use Avogadro's Hypothesis and experimental data to determine the relative mass of molecules.	
2. Use experimental data to determine the relative mass of two objects.	
3. Use experimental data to determine the number of items in a sample without actually counting them.	
4. Given the chemical formula of a substance, determine the molar mass.	
5. Given the mass of a substance, determine <ol style="list-style-type: none"> <li>the number of moles of the sample</li> <li>the number of atoms or molecules in the sample</li> </ol>	

# Chemistry Unit 5 Counting Particles Objective

**Rachel Sandford**



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Fundamentals of Chemistry Fred H. Redmore,1979      **A System of Physical Chemistry ...** William Cudmore  
McCullagh Lewis,1916      Evaluation in a Modular Self-paced Introductory Chemistry Course Eugene Joseph Kales,1978  
*A System of physical chemistry ... v. 1* William Cudmore McCullagh Lewis,1916      *Resources in Education* ,1974  
**Russian Journal of Physical Chemistry** ,1999-10      **Research in Education** ,1974      Teacher's Guidebook for  
Science Problems Wilbur Lee Beauchamp,Joe Young West,1957      **Addison-Wesley Chemistry** ,1990      Psychological  
Monographs Psychological Review Publications,1918      *Energy Research Abstracts* ,1992-03      **Psychological**  
**Monographs** ,1918 Includes music      **The Journal of Industrial and Engineering Chemistry** ,1914      **Webster's New**  
**International Dictionary of the English Language, Based on the International Dictionary 1890 and 1900** William  
Torrey Harris,Frederic Sturges Allen,1911      **Chemical Age** ,1924      **The American Artisan** ,1920      **New**  
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Kaufman Funk,1906      *Nature* Sir Norman Lockyer,1922

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