

Types of Chemical Reactions Worksheet #1

H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
 - 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
 - 3) $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$ **Double Replacement**
 - 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
 - *5) $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
 - 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
 - 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
 - 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
 - 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
 - *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**
- *Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca(OH)}_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe(OH)}_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$ [Copper (II) is used here]

Classifying Chemical Reactions Ch 10 Answer Key

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Prentice Hall Chemistry, 2000 *Introduction to General, Organic, and Biochemistry* Morris Hein, 1997

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Table of Contents Classifying Chemical Reactions Ch 10 Answer Key

1. Understanding the eBook Classifying Chemical Reactions Ch 10 Answer Key
 - The Rise of Digital Reading Classifying Chemical Reactions Ch 10 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Classifying Chemical Reactions Ch 10 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Classifying Chemical Reactions Ch 10 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Classifying Chemical Reactions Ch 10 Answer Key
 - Personalized Recommendations
 - Classifying Chemical Reactions Ch 10 Answer Key User Reviews and Ratings

- Classifying Chemical Reactions Ch 10 Answer Key and Bestseller Lists
- 5. Accessing Classifying Chemical Reactions Ch 10 Answer Key Free and Paid eBooks
 - Classifying Chemical Reactions Ch 10 Answer Key Public Domain eBooks
 - Classifying Chemical Reactions Ch 10 Answer Key eBook Subscription Services
 - Classifying Chemical Reactions Ch 10 Answer Key Budget-Friendly Options
- 6. Navigating Classifying Chemical Reactions Ch 10 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Classifying Chemical Reactions Ch 10 Answer Key Compatibility with Devices
 - Classifying Chemical Reactions Ch 10 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Classifying Chemical Reactions Ch 10 Answer Key
 - Highlighting and Note-Taking Classifying Chemical Reactions Ch 10 Answer Key
 - Interactive Elements Classifying Chemical Reactions Ch 10 Answer Key
- 8. Staying Engaged with Classifying Chemical Reactions Ch 10 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Classifying Chemical Reactions Ch 10 Answer Key
- 9. Balancing eBooks and Physical Books Classifying Chemical Reactions Ch 10 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Classifying Chemical Reactions Ch 10 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Classifying Chemical Reactions Ch 10 Answer Key
 - Setting Reading Goals Classifying Chemical Reactions Ch 10 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Classifying Chemical Reactions Ch 10 Answer Key
 - Fact-Checking eBook Content of Classifying Chemical Reactions Ch 10 Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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