

CLASSIFYING CHEMICAL REACTIONS



Classify the following reactions as synthesis, decomposition, single replacement or double replacement.

ANSWERS

- | | |
|---|--------------------|
| 1. $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ | Decomposition |
| 2. $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$ | Double Replacement |
| 3. $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | Single Replacement |
| 4. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ | Synthesis |
| 5. $2\text{Al} + 3\text{NiBr}_2 \rightarrow 2\text{AlBr}_3 + 3\text{Ni}$ | Single Replacement |
| 6. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ | Synthesis |
| 7. $2\text{NaCl} \rightarrow 2\text{Na} + \text{Cl}_2$ | Decomposition |
| 8. $\text{CaCl}_2 + \text{F}_2 \rightarrow \text{CaF}_2 + \text{Cl}_2$ | Single Replacement |
| 9. $\text{AgNO}_3 + \text{KCl} \rightarrow \text{AgCl} + \text{KNO}_3$ | Double Replacement |
| 10. $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$ | Synthesis |
| 11. $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ | Decomposition |
| 12. $(\text{NH}_4)_2\text{SO}_4 + \text{Ba}(\text{NO}_3)_2 \rightarrow \text{BaSO}_4 + 2\text{NH}_4\text{NO}_3$ | Double Replacement |
| 13. $\text{MgI}_2 + \text{Br}_2 \rightarrow \text{MgBr}_2 + \text{I}_2$ | Single Replacement |
| 14. $\text{SO}_3 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{SO}_4$ | Synthesis |

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