#### COMPUTER SCIENCE PAPER 1 (THEORY)

#### Maximum Marks: 70

#### Time Allowed: Three hours

(Candidates are allowed additional 15 minutes for only reading the paper.

They must NOT start writing during this time).

Answer all questions in Part I (compulsory) and six questions from Part-II, choosing two questions from Section-A, two from Section-B and two from Section-C.

All working, including rough work, should be done on the same sheet as the The intended marks for questions or parts of questions are given in brackets [ ].

#### PART I - 20 MARKS

Answer all questions.

While answering questions in this Part, indicate briefly your working and reasoning, wherever required.

#### Question 1

(i) The law which states  $a + (b.c) = (a+b) \cdot (a+c)$  is:

[1]

- (a) Associative Law
- (b) Distributive Law
- (c) Involution Law
- (d) Commutative Law
- (ii) The dual of  $(P+P) \cdot (Q+0) = Q$  is:

[1]

- (a) P.P' + Q.1 = Q
- (b) PP' + Q0 = Q
- (c) P.P + Q.1 = Q'
- (d) P+P' + Q+1 = Q

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web liquid liquid extraction is an important separation method in research and chemical analysis as a commercial process it is frequently used in the chemical and mining industries and in the downstream recovery of fermentation products antibiotics amino acids steroids

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web in the chemistry lab it is most common to use liquid liquid extraction a process that occurs in a separatory funnel a solution containing dissolved components is placed in the funnel and an immiscible solvent is added resulting in two layers that are shaken together

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web in this work a liquid liquid extraction lle method has been developed to produce liquid pcs via simple mixing and liquid phase separation at room temperature single and two component extractants are used to extract part of the solvent from a diluted colloidal solution which effectively concentrates the colloidal solution without

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web feb 11 2013 summary liquid liquid extraction lle also known as solvent extraction is a typical ternary system in the chemical engineering field which separates chemicals from one solution to another based on the different solubility of the solute chemical in