

Concentration and Molarity PhET Labs

Name: _____

Procedure: https://phet.colorado.edu/sims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all the sliders. Notice what happens to the concentration as solid solute is added and when evaporation occurs.



Concentration

Concentration
(mol/L)
0.179

- How does the concentration change as solid solute is added? concentration increases
 - How does the concentration change as additional water is added? concentration decreases
 - How does the concentration change as evaporation occurs? concentration increases
 - How do you know when a solution is saturated? concentration does not change (it also shows on screen)
 - When a solution is saturated, and additional solid solute is added, what happens? solid does not dissolve and settles at bottom
 - Why do you think this is? solvent can't hold any more solute
 - How does adding this additional solute change the concentration of this saturated solution? decreases concentration
 - How does evaporation change the concentration of a saturated solution? no change- solution is already saturated
- Part 2: Concentrated Solutions (same website as part 1 just click on the Solution option)**
10. Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan here: empty the container and add only one solution from the dropper and then use the sensor to measure concentration

Solution



- Using your plan ... how might you get that concentrated solution to become saturated? you can evaporate
- Does your plan work for all the other solutions too? yes Why? / Why Not? evaporation reduces the amount of water (or solute) so concentration increases



Part 3: Molarity https://phet.colorado.edu/sims/html/molarity/latest/molarity_en.html

Molarity (M) = $\frac{\text{moles of solute (mol)}}{\text{volume of solution (L)}}$

Molarity is moles per liter, that is, how many moles of solute (entire salt) is dissolved per liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Part 5, if your instructor requires it)



Molarity

Cobalt (II) nitrate	5.640 M	Potassium chromate	3.350 M
Cobalt chloride	4.330 M	Nickel (II) chloride	5.210 M
Potassium dichromate	0.510 M	Copper sulfate	1.380 M
Gold (III) chloride	2.30 M	Potassium permanganate	0.480 M

Concentration And Molarity Lab

Judd E. Hollander



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The AGT Cytogenetics Laboratory Manual Marilyn S. Arsham, Margaret J. Barch, Helen J. Lawce, 2017-03-03

Cytogenetics is the study of chromosome morphology structure pathology function and behavior The field has evolved to embrace molecular cytogenetic changes now termed cytogenomics Cytogeneticists utilize an assortment of procedures to investigate the full complement of chromosomes and or a targeted region within a specific chromosome in metaphase or interphase Tools include routine analysis of G banded chromosomes specialized stains that address specific chromosomal structures and molecular probes such as fluorescence in situ hybridization FISH and chromosome microarray analysis which employ a variety of methods to highlight a region as small as a single specific genetic sequence under investigation The AGT Cytogenetics Laboratory Manual Fourth Edition offers a comprehensive description of the diagnostic tests offered by the clinical laboratory and explains the science behind them One of the most valuable assets is its rich compilation of laboratory tested protocols currently being used in leading laboratories along with practical advice for nearly every area of interest to cytogeneticists In addition to covering essential topics that have been the backbone of cytogenetics for over 60 years such as the basic components of a cell use of a microscope human tissue processing for cytogenetic analysis prenatal constitutional and neoplastic laboratory safety and the mechanisms behind chromosome rearrangement and aneuploidy this edition introduces new and expanded chapters by experts in the field Some of these new topics include a unique collection of chromosome heteromorphisms clinical examples of genomic imprinting an example driven overview of chromosomal microarray mathematics specifically geared for the cytogeneticist usage of ISCN s cytogenetic language to describe chromosome changes tips for laboratory management examples of laboratory information systems a collection of internet and library resources and a special chapter on animal chromosomes for the research and zoo cytogeneticist The range of topics is thus broad yet comprehensive offering the student a resource that teaches the procedures performed in the cytogenetics laboratory environment and the laboratory professional with a peer reviewed reference that explores the basis of each of these procedures This makes it a useful resource for researchers clinicians and lab professionals as well as students in a university or medical school setting

Essential Laboratory Mathematics Catherine W. Johnson, Daniel L. Timmons, Pamela E. Hall, 2009-12-03 This hands on manual with pedagogical features that draw the learner into the content offers clear and complete coverage of the mathematical topics most often used in today s clinical and medical laboratories Furthermore it provides a solid foundation for subsequent courses in the laboratory sciences The first two chapters present a review of basic mathematical concepts The remainder of the book provides students with a realistic means to build on previously learned concepts both mathematical and scientific to refine their mathematical skills and to gauge their mastery of those skills Outstanding features Each chapter opens with an outline objectives and key terms Key terms highlighted within the text are listed and defined in the glossary Margin problems and practice problem sets provide the chance to gain immediate

proficiency Laboratory exercises and review problems allow students to apply what they've learned and assess their understanding and progress A special calculator icon signals explanations of calculator use for a particular mathematical function Study hints Keys to Success offer practical suggestions and guidance for maximizing achievement The workbook design enables users to solve problems and take notes directly on the pages

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Virtual Chemlab Brian F. Woodfield, Matthew C. Asplund, Steven Haderlie, 2006 Contains a full virtual lab environment as well as the pre arranged labs that are referenced in the workbook and at the end of the chapter in the textbook Virtual ChemLab can be run directly from the CD or installed on the student's computer **Exploring Chemistry Laboratory Experiments in General, Organic and Biological Chemistry** Julie R. Peller, 2003-04 This lab manual is organized and written to ensure that non science majors are comfortable with chemistry labs by making the experiments more applicable to students daily lives This approach also serves to make the experiments more understandable Many labs relate specifically to allied health fields PE Lab Exp(Noncons)Mod Chem 90 Tzimopoulos, 1990 **Elements of General and Biological Chemistry** John R. Holum, 1987 This book is a shorter version of the third edition of Fundamentals of General Organic and Biological Chemistry 1986 It incorporates the recommendations of the Task Force on Chemical Education for Health Professions and meets the needs for a basic text in a one term course in chemistry for students aiming for careers in professional health care fields *Lab Experiments for Modern Chemistry* Tzimopoulos, 1990 Laboratory Experiments for General Chemistry Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998 *Merrill Chemistry-Lab.Manual* Smoot, 1994-07 Lab Experiments Modern Chemistry Harold Clark Metcalfe, 1982 **Journal of Clinical & Laboratory Immunology**, 1984 The Extraordinary Chemistry of Ordinary Things, Study Guide Carl H. Snyder, David F. Dever, 1995-01-19 Examines the chemistry of the substances of our everyday world Our daily lives are immersed in chemicals an effective way to teach and learn chemistry is by examining the goods and substances that we use in our daily lives and that affect us and our environment **Laboratory Exercises in Biology** Theodore T. Ziegenfuss, 1990 Instruments & Control Systems, 1974

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