

## Study Card: Solvation

### What Does Dissolving Involve?

When a solid dissolves in water, it's not *gone*. It has changed its state. As a solid, a solute particle (ion, molecule, etc.) is attracted to other solute particles. But when it dissolves in  $\text{H}_2\text{O}$ , it changes to the aqueous state and becomes attracted to the  $\text{H}_2\text{O}$  molecules.

### The Dissolving of Ionic Solids

Dissolving involves more than just *mixing*. When table salt ( $\text{NaCl}$ ) is added to water, the ions break off the crystal one-by-one and are surrounded by  $\text{H}_2\text{O}$  molecules. The O atom of  $\text{H}_2\text{O}$  surrounds the + ions. The H atoms of  $\text{H}_2\text{O}$  surround the - ions. This is known as **solvation**; it prevents the ions from returning to the crystal lattice. All ionic solids are dissolved in this manner.

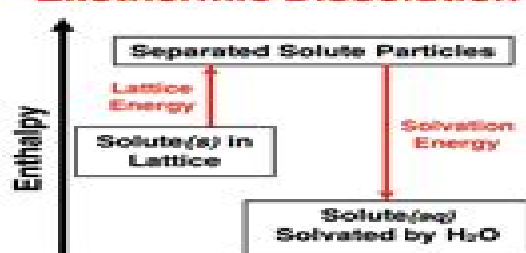
### The Dissolving of Molecular Solids

Molecular solids consist of molecules instead of ions. When table sugar (sucrose,  $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ) dissolves, the molecules break off the crystal one-by-one and are surrounded by  $\text{H}_2\text{O}$  molecules. Hydrogen bonding forces hold the sucrose molecules together. So  $\text{H}_2\text{O}$  surrounds the -OH groups of a sucrose molecule, forming its own hydrogen bonds with sucrose and preventing the molecule from returning to the crystal. This is how **solvation** helps molecular solids dissolve.

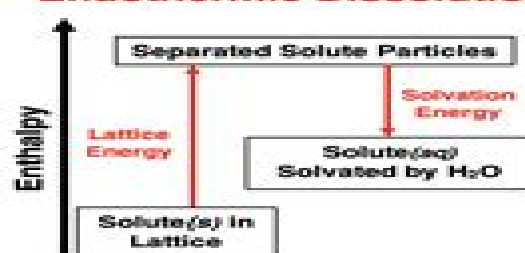
### An Energy Perspective on Dissolving

Dissolving can be thought of as involving two steps. First, solute particles are separated from the crystal. This requires energy; the amount is roughly equal to the **lattice energy**. Second, the  $\text{H}_2\text{O}$  molecules surround the solute particles. This releases energy; the amount is roughly equal to the **solvation energy**. Dissolution is more likely to occur if the solvation energy is greater than the lattice energy.

#### Exothermic Dissolution



#### Endothermic Dissolution



## The Dissolving Process



Solute particles  
attract each other.

Solute particles  
attract water.

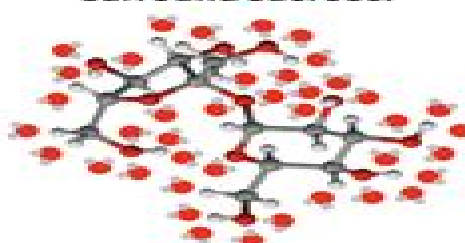
### Solvation

Polar  $\text{H}_2\text{O}$  molecules surrounding + and - ions.



### Solvation

$\text{H}_2\text{O}$  molecules attract and surround sucrose.



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# Concept Review How Substances Dissolve Answers

**Yijin Wang**



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## **Table of Contents Concept Review How Substances Dissolve Answers**

1. Understanding the eBook Concept Review How Substances Dissolve Answers
  - The Rise of Digital Reading Concept Review How Substances Dissolve Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Concept Review How Substances Dissolve Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Concept Review How Substances Dissolve Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Review How Substances Dissolve Answers
  - Personalized Recommendations
  - Concept Review How Substances Dissolve Answers User Reviews and Ratings
  - Concept Review How Substances Dissolve Answers and Bestseller Lists
5. Accessing Concept Review How Substances Dissolve Answers Free and Paid eBooks
  - Concept Review How Substances Dissolve Answers Public Domain eBooks
  - Concept Review How Substances Dissolve Answers eBook Subscription Services
  - Concept Review How Substances Dissolve Answers Budget-Friendly Options
6. Navigating Concept Review How Substances Dissolve Answers eBook Formats

- ePub, PDF, MOBI, and More
- Concept Review How Substances Dissolve Answers Compatibility with Devices
- Concept Review How Substances Dissolve Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Concept Review How Substances Dissolve Answers
  - Highlighting and Note-Taking Concept Review How Substances Dissolve Answers
  - Interactive Elements Concept Review How Substances Dissolve Answers
- 8. Staying Engaged with Concept Review How Substances Dissolve Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Concept Review How Substances Dissolve Answers
- 9. Balancing eBooks and Physical Books Concept Review How Substances Dissolve Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Concept Review How Substances Dissolve Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Concept Review How Substances Dissolve Answers
  - Setting Reading Goals Concept Review How Substances Dissolve Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concept Review How Substances Dissolve Answers
  - Fact-Checking eBook Content of Concept Review How Substances Dissolve Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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