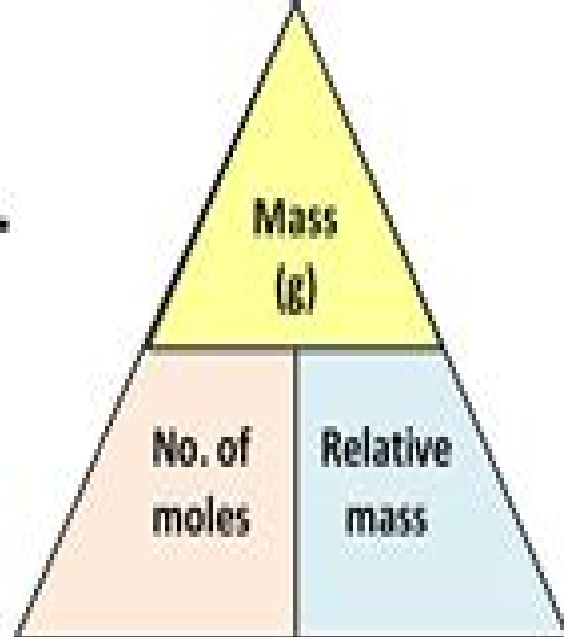


# Calculating limiting reactants.

Let's look at an example:



In this reaction, 56 g nitrogen gas ( $\text{N}_2$ ) was added to a chamber containing 14 g hydrogen gas ( $\text{H}_2$ ). Which was the limiting reactant? ( $A_r$ :  $\text{H}=1$ ,  $\text{N}=14$ )

First steps?

# Chemistry Unit 8 Limiting Reactant

**Daniele Peters**



## **Chemistry Unit 8 Limiting Reactant:**

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Study Guide General Chemistry, Second Edition, Becker/Wentworth Becker, Benjamin L. Carroll, 1980

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