

→ HEAT

Chapter - 11

CALORIC THEORY:

In 18th Century, physicist believed that heat in a weightless, invisible and indestructible fluid called "Caloric". It was assumed to exist in every material body. Hot bodies contain more caloric than cold bodies. This explain the conduction and mixing of substances.

MODERN CONCEPT:

This concept of heat fluid was challenged by count Rumford. He observed while supervising the boring of cannon that cannon barrels were being bored a tremendous amount of heat was given off. In order to further investigate it, Rumford made use of a very dull drill, which was unable to cut the metal. Heat was evolved in apparently unlimited quantity. He conclude that heat was due to the rotation of the borer and not from the metal itself. As a result of scientists came to interpret.

"Heat a transient flow that transfers from hot body to cold body due to difference of temperature".
Once heat is transferred into the body it is converted into internal energy of the system.

Units:

S.I unit of heat is Joule (J).

Temperature:

It is defined as "Measure of degree of hotness or coldness of a body on some chosen scale".

OR

It is also defined as: on microscopic level as:

"Temperature is a measure of the average translational kinetic energy of the molecules of a body".

SCALES OF TEMPERATURE:

i. Celsius Scale (Centigrade Scale):

The lower fixed point called ice point is the melting point of ice (0°C), upper fixed point called steam point is the boiling point of water, (100°C), at standard pressure. It is divided into 100 equal intervals.

ii. Fahrenheit Scale:

The lower fixed point is 32 and upper fixed point is 212 and the interval between them is equally divided into 180 equal parts.

iii. Kelvin Scale (Absolute Scale):

The thermodynamic temperature is measured in Kelvin. It is SI unit of temperature. This scale uses one fixed point called "Triple point of water". The lowest temperature on this scale is " -273°C ".

$^{\circ}\text{C}$	$^{\circ}\text{F}$	K
100	212	373
0	32	273
-273	460	0

Class 12 Physics Heat Notes

Didier Musso



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