

Addition Compounds

Double Salt.

Complex Salt.

* Carnallite :-



* Mohr's Salt :-



* Potash Alum :-



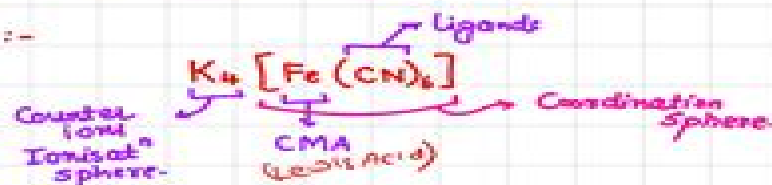
* d-block elements.

↳ * Small Size,

* high ϕ Nature.

* Vacant Orbitals

Note :-



* Denticity :- No. of e^- pairs which can be donated by a ligand to metal ion.

* In any ligand if total no. of atoms are 3 or less than 3 $\rightarrow$ then denticity is always "1".

Stable $\rightarrow$ free

CLASSIFICATION OF LIGANDS :-

A) Anionic

SO_4^{2-} = Sulfate

CO_3^{2-} = Carbonate

Cl^- = Chloride

OH^- = Hydroxide

F^- = Fluoride

NH_2^- = Amide

NH^{2-} = Imide

N_3^- = Azide

N^{3-} = Nitride

O^{2-} = Oxide

O_2^- = Superoxide

O_2^{2-} = Peroxide

CN^- = Cyanide

NC^- = Isocyanide

NO^- = Nitrite-N

ONO^- = Nitrite-O

OCN^- = Cyanate-O

NCO^- = Cyanate-N

SCN^- = Thiocyanate-S

NCS^- = Thiocyanate-N

$C_2O_4^{2-}$ = Oxalate (2)

Gly = Glycinate (2)

dimethylgly = dimethylglycinate (2)

glyoximate (2)

$EDTA^{4-}$ = Ethylene diamine tetra-acetate. (6)

B) Neutral :-

H_2O :- Aqua

NH_3 :- Amine

CO :- Carbonyl

NO :- Nitrosyl

PH_3 :- Phosphine

PPh_3 :- Triphenyl Phosphine

R_2 :- Pyridine.

en :- ethylene diamine.

pn :- propylene diamine

tn :- trimethylene diamine

dien :- diethylene triamine

taien :- triethylene tetraamine.

C) Cationic :-

NO^+ = Nitrosonium

$NH_2-NH_2^+$:- Hydrazonium

Jugaad :-

$\rightarrow Pt/Pd$

always dsp^2

$CN=4$

O.S. = +2

$CN=6$

O.S. = +4.

$\rightarrow$ When Fe & NO both +1.

$NO = +1$

$\rightarrow Sc Ti V Cr Mn Fe Co Ni Cu Zn \}^{+2}$

3d = 1 2 3 4 5 6 7 8 9 10

Similar for below 4d & 5d !

$\rightarrow d^1/d^2/d^3 \xrightarrow[6 \text{ ligands}]{sp^3d^2} d^2sp^3$

UPe :- 1 2 3 $CN=6$ always Ioc LSC Paramag.

$\rightarrow d^4/d^5/d^6 \xrightarrow[6 \text{ ligands}]{sp^3d^2} sp^3d^2$

UPe :- 2 1 0 $CN=6$ always ooc HSC

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