

## COORDINATION COMPOUND

### # Introduction:-



#### Double salt

- Completely ionise in aq. medium.
- Lose their identity in aq. soln.
- Each ion gives positive test.
- Individual ions do not lose their own identity.

#### Examples

- ① Carnallite  $[KCl \cdot MgCl_2 \cdot 6H_2O]$
- ② Mohar's salt  $[FeSO_4(NH_4)_2SO_4 \cdot 6H_2O]$
- ③ Potash alum  $[K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O]$

#### Complex salt

- Do not ionise completely.
- Retain their identity in aq. medium.
- Each ion do not give positive test.
- Ions lose their own identity.

Examples:-  $K_4[Fe(CN)_6]$ ,  $[Pt(NH_3)_2Cl_2]$  etc.

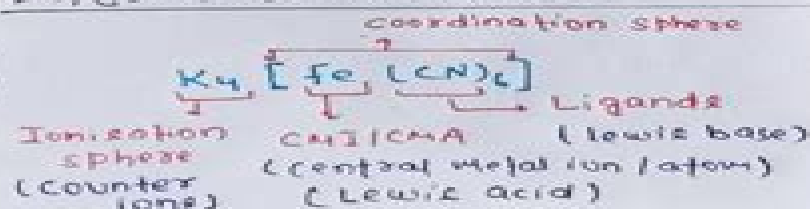
$\hookrightarrow$  Complex salt  $\rightarrow$   $\sigma$  coordinate bonds present b/w ligand ( $e^-$  pair donors) & metal ion ( $e^-$  pair acceptor).

Q  $\rightarrow$  Why d-block element (Metal) form more no. of complexes?

Reason:- ① Availability of vacant orbital

② Small size & High  $\phi$  nature

$\phi \uparrow$ , Covalent nature of Bond  $\uparrow$



## # Definitions

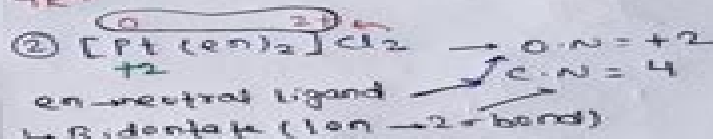
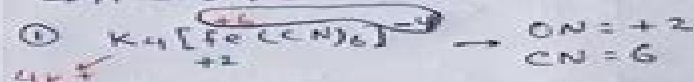
① Coordination number:- No. of ligand donor atoms to which metal is directly bonded. or

$\hookrightarrow$  coordination no. is equal to  $\sigma$ -bonds present b/w ligand & CMI.

According to NCERT,  $\pi$ -bond b/w CMI & ligands will not be included in CN.

② oxidation no. of central atom:-

$\hookrightarrow$  Andaz & Bahar ke charge ka hi saaf



③ Homoleptic complexes:- CMI is bound to only one kind of ligands.



④ Heteroleptic complexes:- CMI is bound to more than one kind of ligands.  
 Ex  $\rightarrow [Pt(NH_3)_2Cl_2]$

⑤ Coordination polyhedron:- Dabbe ki geometry (decided by CN of complex)

(C.N)

2  $\rightarrow$  Linear

3  $\rightarrow$  T planar

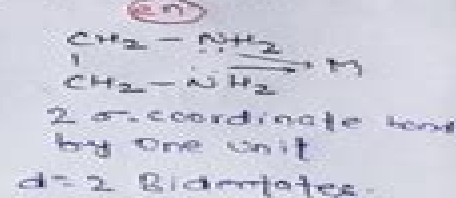
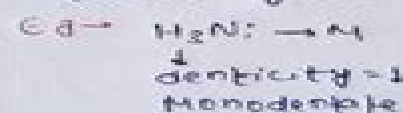
4  $\rightarrow$  Sq planar

\* 4  $\rightarrow$  Tetrahedral

5  $\rightarrow$  TBP

\* 6  $\rightarrow$  S.B.P or octahedral

# Denticity:- no. of  $\sigma$ -coordinate bond formed by one unit of ligand.



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1. Understanding the eBook Class 11 Coordination Compounds Notes From
  - The Rise of Digital Reading Class 11 Coordination Compounds Notes From
  - Advantages of eBooks Over Traditional Books
2. Identifying Class 11 Coordination Compounds Notes From
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Class 11 Coordination Compounds Notes From
  - User-Friendly Interface
4. Exploring eBook Recommendations from Class 11 Coordination Compounds Notes From
  - Personalized Recommendations
  - Class 11 Coordination Compounds Notes From User Reviews and Ratings

- Class 11 Coordination Compounds Notes From and Bestseller Lists
- 5. Accessing Class 11 Coordination Compounds Notes From Free and Paid eBooks
  - Class 11 Coordination Compounds Notes From Public Domain eBooks
  - Class 11 Coordination Compounds Notes From eBook Subscription Services
  - Class 11 Coordination Compounds Notes From Budget-Friendly Options
- 6. Navigating Class 11 Coordination Compounds Notes From eBook Formats
  - ePub, PDF, MOBI, and More
  - Class 11 Coordination Compounds Notes From Compatibility with Devices
  - Class 11 Coordination Compounds Notes From Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Class 11 Coordination Compounds Notes From
  - Highlighting and Note-Taking Class 11 Coordination Compounds Notes From
  - Interactive Elements Class 11 Coordination Compounds Notes From
- 8. Staying Engaged with Class 11 Coordination Compounds Notes From
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Class 11 Coordination Compounds Notes From
- 9. Balancing eBooks and Physical Books Class 11 Coordination Compounds Notes From
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Class 11 Coordination Compounds Notes From
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Class 11 Coordination Compounds Notes From
  - Setting Reading Goals Class 11 Coordination Compounds Notes From
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Class 11 Coordination Compounds Notes From
  - Fact-Checking eBook Content of Class 11 Coordination Compounds Notes From
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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