

Practice makes perfect with Faraday's Law
Using conversion factors in Electrochemistry

1. Determine the mass of copper that is plated out when a current of 10.0 amps is passed for 30.0 minutes through a solution containing Cu^{2+}

$$\text{Cu}^{+2} + 2\text{e}^- \rightarrow \text{Cu}(s)$$

$$30 \text{ min} \times \frac{60 \text{ s}}{1 \text{ min}} = 1800 \text{ s} \quad / \quad \frac{1 \text{ mole e}^-}{96485 \text{ C}} \times \frac{1 \text{ mole Cu}}{2 \text{ mole e}^-} \times \frac{63.55 \text{ g}}{1 \text{ mole Cu}}$$

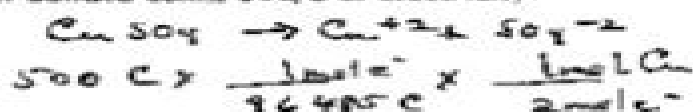
5.83 g Answer

2. How long will it take to plate out 1.0 kg Al from aqueous Al using a current of 100.0 A? Calculate in minutes

$$1000 \text{ g Al} \times \frac{1 \text{ mole Al}}{26.98 \text{ g}} \times \frac{3 \text{ mole e}^-}{1 \text{ mole Al}} \times \frac{96485 \text{ C}}{1 \text{ mole e}^-} \times \frac{1 \text{ sec}}{1000 \text{ C}} \times \frac{1 \text{ min}}{60 \text{ s}}$$

1788 min Answer 1800 min

3. Calculate the moles of copper metal that can be produced by the electrolysis of molten copper sulfate using 500 C of electricity.



0.0259 or 2.59 x 10⁻³ Answer moles

4. What mass of copper could be deposited from a copper(II) sulfate solution using a current of 0.50 A over 10 seconds?

$$\text{CuSO}_4 \rightarrow \text{Cu}^{+2} + \text{SO}_4^{-2}$$

$$.50 \frac{\text{C}}{\text{s}} \times 10 \text{ s} \times \frac{1 \text{ mole e}^-}{96485 \text{ C}} \times \frac{1 \text{ mole Cu}}{2 \text{ mole e}^-} \times \frac{63.55 \text{ g}}{1 \text{ mole Cu}}$$

1.6 x 10⁻³ g Answer

5. Calculate the time required to deposit 56 g of silver from a silver nitrate solution using a current of 4.5 A.



11631 s = 11000 s Answer
3.1 hrs

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Puja Mehta



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