# Calculating Limiting Reactant, Theoretical Yield, and Percent Yield from Reactant Masses

$$\begin{array}{c} \textbf{42.5 g Mg} \times \frac{1 \, \text{mol-Mg}}{24.31 \, \text{g Mg}} \times \frac{2 \, \text{mol MgO}}{2 \, \text{mol-Mg}} = \begin{array}{c} \textbf{1.7483 \, mol MgO} \\ \textbf{Limiting} \\ \text{reactant} \end{array} \\ \textbf{33.8 g } \Theta_2^- \times \frac{1 \, \text{mol-O}_2^-}{32.00 \, \text{g} \cdot \Theta_2^-} \times \frac{2 \, \text{mol MgO}}{1 \, \text{mol-O}_2^-} = \begin{array}{c} \textbf{2.1125 \, mol MgO} \\ \textbf{2.1125 \, mol MgO} \end{array} \end{array} \right) \\ \textbf{1.7483 \, mol MgO} \times \frac{40.31 \, \text{g MgO}}{1 \, \text{mol-MgO}} = 70.5 \, \text{g MgO}$$

# **Concept Review Limiting Reactant Percent Yield**

**Charles H. Corwin** 

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Holt Chemistry Holt Rinehart & Winston, 2003-01-24

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